

Proceeding Paper

Synthetic Access to Tetracationic Benzoporphyrins and Their Role as Photosensitizers towards Gram-Negative *Escherichia coli* †

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† Presented at the 26th International Electronic Conference on Synthetic Organic Chemistry; Available online: <https://ecsoc-26.sciforum.net>.

Abstract: Nowadays world population faces an episode where bacteria are becoming resistant to antibiotics, and it is crucial to find out alternatives and new molecules to fight these microorganisms. Photodynamic inactivation of microorganisms has been pointed out as an alternative to conventional therapies. This work describes the synthesis of benzoporphyrins derivatives bearing triazolyl groups and of the analogues with pyridyl units under Heck coupling conditions. The benzoporphyrin derivatives containing pyridyl groups were further quaternized with iodomethane and 1-iodopentane to evaluate the influence of alkyl chain size on their photoinactivation ability. The biological studies towards Gram-negative bioluminescent *Escherichia coli* showed that the tetracationic benzoporphyrins can efficiently inactivate this bacterium.

Keywords: Tetracationic benzoporphyrins; Heck coupling reaction; Antimicrobial Photodynamic Therapy; Bioluminescent *E. coli*

Citation: Morais, F.M.P.; Vieira, C.; Gomes, A.T.P.C.; Faustino, M.A.F.; Almeida, A.; Neves, M.G.P.M.S.; Moura, N.M.M. Synthetic Access to Tetracationic Benzoporphyrins and Their Role as Photosensitizers towards Gram-Negative *Escherichia coli*. *Chem. Proc.* **2022**, *4*, x. <https://doi.org/10.3390/xxxxx>

Academic Editor(s): Julio A. Seijas

Published: 15 November 2022

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1. Introduction

Porphyrins are tetrapyrrolic macrocycles with important roles in some biological processes like respiration or electron transportation. Due to their unique physicochemical properties, porphyrins are currently being used in a wide range of applications like artificial photosynthesis, catalysis, sensors, nonlinear optics, and medicine [1,2]. In this specific field, considerable attention is being given to their use as photosensitizers (PS) in photodynamic therapy (PDT) of cancer [3,4] or, more recently, in the photoinactivation of microorganisms usually referred to as antimicrobial photodynamic therapy (aPDT) [5,6].

Like in PDT, aPDT requires the presence of a PS, light and dioxygen (³O₂). The association of these components results in the formation of reactive oxygen species (ROS), such as singlet oxygen (¹O₂) that cause damage to vital biomolecules leading to microbial death [7–9]. The multitarget action of aPDT supports its selection as an excellent alternative to conventional antimicrobial drugs considering the emergence of microbial resistance to many of those therapeutic agents used in clinical and environmental contexts [10]. It has been shown that the type of bacterial membrane and the PS ability to bind to it are factors that influence antimicrobial photoinactivation [8].

Due to the chemical and photophysical properties of porphyrins, this family of compounds is considered the foremost used as PSs in aPDT [11]. Therefore, it is crucial to develop novel synthetic approaches or to improve the outcomes of the ones already described

to provide porphyrin derivatives with suitable characteristics for their application in aPDT. In this work, we report the preparation under conventional and microwave (MW) heating of benzoporphyrins derivatives bearing triazolyl groups and of the analogues with pyridyl units under Heck coupling conditions. The neutral pyridyl derivatives were then submitted to cationization by using appropriate alkylating agents.

2. Experimental

2.1. Synthesis of the Starting Porphyrins and Benzoporphyrin 2a

Starting porphyrins **1a,b** and benzoporphyrin derivative **2a** were synthesized following well-established procedures already reported in the literature. The structure of these compounds was confirmed by ¹H NMR spectroscopy and mass spectrometry and agrees with the analytical data described in the literature [12–15].

2.2. Synthesis of Benzoporphyrin 2b

2.2.1. Under Conventional Heating Conditions

In a 25 mL flask was added the nickel(II) complex **1b** (20.3 mg, 0.020 mmol), Pd(OAc)₂ (2.3 mg, 0.010 mmol), PPh₃ (9.9 mg, 0.038 mmol), K₂CO₃ (12.2 mg, 0.089 mmol), 1-phenyl-4-vinyl-1,2,3-triazole (28.0 mg, 0.164 mmol). The solids, under a nitrogen atmosphere, were dissolved in toluene (1 mL) and DMF (0.5 mL). The reaction was kept at 120 °C for 6 h when the thin-layer chromatography (TLC) control showed the total consumption of the starting Ni(II) porphyrin derivative **1b**. The mixture was washed with water and the organic phase was extracted with toluene. The crude was purified by preparative TLC using CHCl₃ and MeOH (4:1) as eluent. Triazolo-benzoporphyrin **2b** was obtained in 15% yield (3.1 mg).

2.2.2. Under Microwave Irradiation

The same amounts of nickel(II) complex **1b**, triazole and catalytic system used in the conventional heating method were added to a microwave reactor and dissolved in a 2:1 toluene/DMF mixture (1.5 mL). The reaction mixture was subjected to microwave radiation for 5 min at 120 °C, at a power of 250 W in a CEM Microwave Synthesis System Discover. At the end of this time, it was found that all the starting material **1b** had been consumed. The reaction was washed with water and the organic phase was extracted with toluene. The reaction mixture was purified by preparative thin layer chromatography using first CHCl₃ as eluent and then a mixture of CHCl₃/MeOH (4:1). The desired benzoporphyrin **2b** bearing triazolyl units was obtained in 10% yield (2.0 mg).

Yield: 10%. ¹H NMR (300 MHz, CDCl₃): δ 8.73 (s, 4H, β-H), 8.03–8.00 (m, 12H, H-*o*-Ph and H-5'-Triazole), 7.80–7.77 (m, 20H, H-*o*-Ph-Triazole and H-*m,p*-Ph); 7.59 (t, *J* = 7.6 Hz, 8H, H-*m*-Ph-Triazole), 7.49 (d, *J* = 7.6 Hz, 4H, H-*p*-Ph-Triazole), 7.44 (s, 4H, H-1'). **MS-ESI(+):** *m/z* calculated for C₈₄H₅₂N₁₆Ni [M]⁺• 1344.2; found 1344.0.

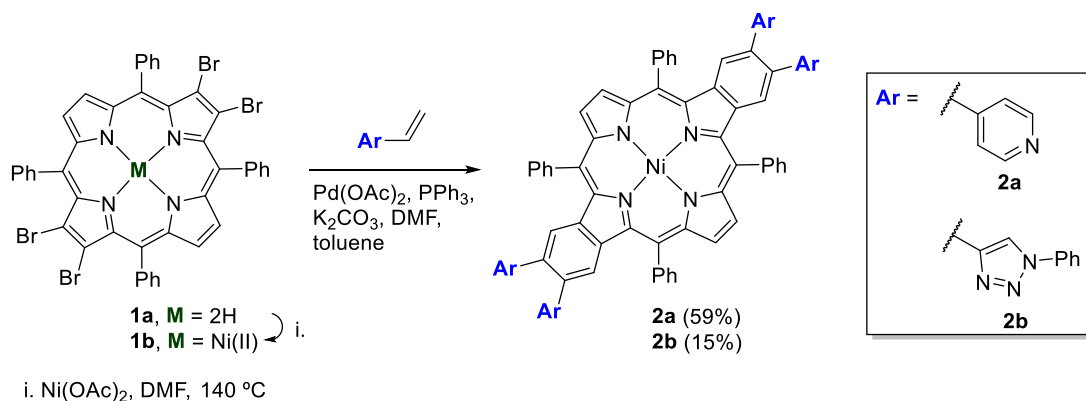
2.3. Synthesis of the Cationic Benzoporphyrins 4a,b

The charged benzoporphyrins **4a,b** were prepared from the free-base derivative **3** using methyl iodide or 1-iodopentane as alkylating agents following well-established procedures already reported in the literature [16]. The structure of these compounds was confirmed by ¹H NMR spectroscopy and mass spectrometry and is in agreement with the described in the literature [13].

3. Results and Discussion

3.1. Synthesis

The synthetic route to access the tetracationic triazolyl and pyridyl porphyrins required the preparation of the 5,10,15,20-tetraphenylporphyrin (TPP) for further bromination at beta pyrrolic positions to attain compound **1a**. These starting porphyrins were prepared following a well-established procedure reported in the literature [13].



Scheme 1. Heck reaction conditions used to synthesize benzoporphyrin derivatives **2a,b**.

Once the Heck reaction is catalyzed by palladium it was necessary to protect the core of the macrocycle in order to minimize the loss of the catalyst during the catalyst cycle through the coordination of the porphyrin with Pd(II) [17]. Thus, porphyrin **1a** was complexed with Ni(II) using nickel(II) acetate in DMF under reflux for 1 h. After the reaction was completed, Ni(II) porphyrinate **1b** was obtained with a yield of 67%. The simplification of the UV-Vis spectrum of **1b** in the Q band region, when compared with the one of porphyrin **1a** (two bands versus four bands) confirmed the success of this complexation step. The first assays concerning the Heck coupling reaction between the nickel(II) porphyrinate **1b** and 4-vinylpyridine in the presence of Pd(OAc)₂, PPh₃ and K₂CO₃ were performed in a mixture of dimethylformamide (DMF) and toluene under conditions like those described by Jiang et al. [15]. After 6 h at 120 °C, the TLC monitorization confirmed the full consumption of the starting material (compound **1b**) and the formation of several more polar compounds. After work-up and preparative chromatography, the most polar and main compound was identified as the desired derivative **2a** and was isolated in 15% yield. In order to minimize the formation of secondary products and to improve the yield of compound **2a**, the Heck coupling reaction was repeated but now under microwave irradiation. This experiment was performed in a MW reactor, for 5 min at 120 °C and at a power of 250 W and using the same catalytic conditions previously reported to couple nickel(II) porphyrinate and 4-vinylpyridine. Benzoporphyrin **2a** was obtained with a yield of 59%, substantially higher than that obtained in the reaction performed under conventional heating conditions.

Due to the success achieved in the synthesis of benzoporphyrin derivative **2a**, this methodology was extended to the reaction between Ni(II) porphyrinate **1b** and 1-phenyl-4-vinyl-1,2,3-triazole. Triazoles play an important role in medicinal and organic chemistry due to their fairly simple synthesis, versatility and remarkable biological activity [18,19]. Firstly, we decided to try the conventional heating approach used by Jiang et al. [15]. procedure reported for benzoporphyrin **2a**. Also, in this case, the most polar and main product was identified as the desired benzoporphyrin bearing the four triazolyl units and was isolated in 15% yield; unfortunately, many other compounds were obtained from the reaction in amounts that did not allow their identification.

With the main purpose to increase the yield of the reaction and due to the previous success, we decided to perform an optimization study under MW irradiation conditions.

In this study, were evaluated several parameters such as time and addition of an oxidizing agent (nitrobenzene) at different reaction temperatures. The different conditions are summarized in Table 1 and the results obtained allowed us to conclude that the better conditions to obtain **2b** are under conventional heating conditions followed by MW irradiation for 10 min at 120 °C and in the absence of the oxidant.

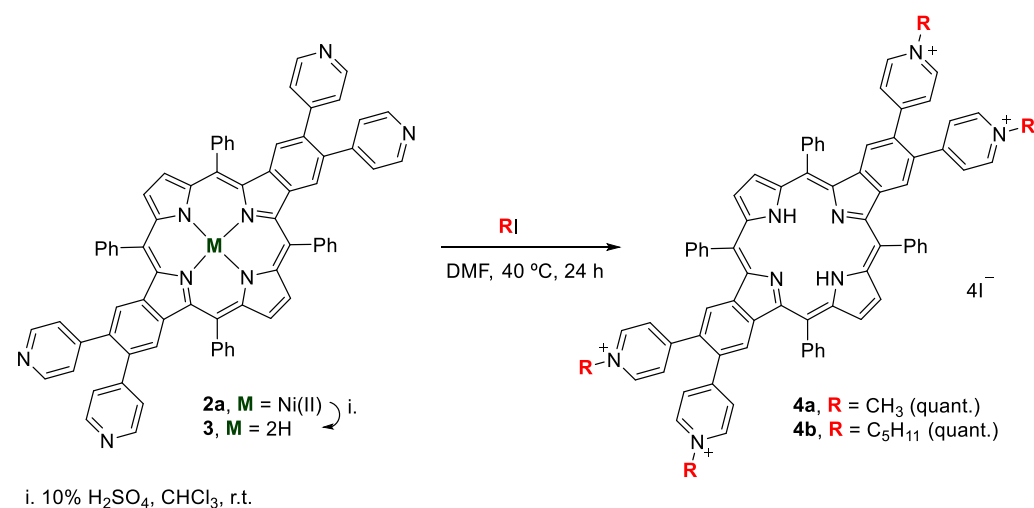
Table 1. Conditions studied aiming to prepare benzoporphyrin derivatives **2b**.

Entry	Heating Source	Toluene/DMF Ratio	Oxidant	T (°C)	Time	Yield (%)
1	Oil bath	2:1	-	120	6 h	15
2			-		10 min	10
3	-		15 min		trace	
4	MW		-		15 min	trace
5			C ₆ H ₅ NO ₂		10 min	trace
6			C ₆ H ₅ NO ₂		140	10 min

Due to the very low yield obtained for triazolyl benzoporphyrin derivative **2b**, we decided not to proceed with this compound in the following synthetic steps, as well as, with the biological evaluation.

Knowing that the Ni(II) complexes of porphyrins present poor efficiency to generate singlet oxygen, but the presence of positive charges improves the photodynamic inactivation [20], the benzoporphyrin **2a** was first demetallated for further cationization.

In order to evaluate the influence of alkyl chain size on singlet oxygen generation and on the photoinactivation of antibiotic-resistant bacteria, the cationization of the pyridyl groups of the benzoporphyrin derivative **3** with iodomethane and 1-iodopentane was performed (Scheme 2); the desired cationic porphyrins **4a** and **4b** were obtained quantitatively.



Scheme 2. Synthetic approach to attain cationic benzoporphyrin derivatives **4a,b**.

3.2. Photophysical and Photochemical Properties and Photoinactivation of Bioluminescent *E. coli*

The UV-Vis absorption spectrum was recorded for each benzoporphyrin. It was possible to conclude that the modifications in the benzoporphyrin core structure influence the absorption spectrum of the different compounds. The electronic spectrum of Ni(II) complex **2a** shows the typical sharp band at 444 nm (Soret band) and two less intense bands but well-defined bands at 559 and 598 nm, respectively. The corresponding free-base derivative **3** displays the Soret band at 436 nm and four Q bands ranging from 526 to 670 nm. After

quaternization of the pyridyl units, the Soret band of each compound **4a** and **4b** are red-shifted ca. 30 nm and is possible to observe three Q-bands from 536 to 613 nm.

The evaluation of the stability of a molecule under irradiation is mandatory before evaluating its efficiency to generate singlet oxygen generation. Therefore, benzoporphyrin **3** and tetracationic benzoporphyrins derivatives **4a** and **4b** were exposed to white light irradiation for 30 min with a power of 11 mW/cm². The absence of any significant change in their UV-Vis spectra confirmed that these compounds are photostable.

One of the most relevant features of a PS is its aptitude to generate singlet oxygen which is strongly related to the efficiency of the photodynamic process [20]. The capability of the PSs synthesized to generate ¹O₂ was quantitatively assessed by monitoring the decomposition of 1,3-diphenylisobenzofuran (DPiBF) at 415 nm in DMF/H₂O (9:1). Therefore, the absorbance decrease of DPiBF allowed us to evaluate the generation of ¹O₂ by the PS present in solution [21].

The data obtained with neutral benzoporphyrin derivative **3** and tetracationic benzoporphyrins derivatives **4a** and **4b** (Figure 1) confirmed their ability to induce the decay of DPiBF absorption when irradiated with red light in the presence of each PS. In addition, tetracationic benzoporphyrin derivative **4b** showed a superior ¹O₂ production capacity than Tetra-Py(+)-Me, a reference compound pointed out as a good ¹O₂ generator [22].

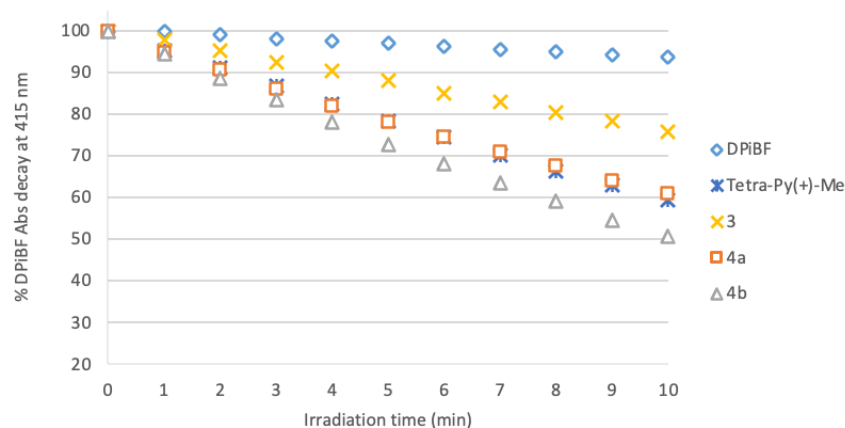


Figure 1. Time-dependent photodecomposition of DPiBF at 50 μ M photosensitized by derivatives **3**, **4a,b** and Tetra-Py(+)-Me at 0.5 μ M in DMF/H₂O (9:1) upon irradiation with red light (630 ± 20 nm) at an irradiation of 11 mW/cm².

Preliminary studies show that both tetracationic benzoporphyrin derivatives **4a** and **4b** display the capability to act as a PS in the photoinactivation of bioluminescent gram-negative *E. coli*. The pentylated derivative showed to be more efficient in *E. coli* photoinactivation when compared to the methylated porphyrin, reaching the detection limit of the method after 180 min of irradiation, probably due to its larger alkyl chain, which promotes better bacterial membrane binding.

4. Conclusions

This work describes a synthetic approach that allowed for the first time to prepare a benzoporphyrin derivative bearing triazolyl moieties via Heck reaction coupling reaction. However, the low yield obtained makes mandatory further synthetic studies to improve the yield of this derivative. The yield of the benzoporphyrin derivative bearing pyridyl units at beta-pyrrolic positions was improved under microwave irradiation and the products obtained after quaternization of the pyridyl units with appropriate alkylating agents, display the requirements and potential to be used as photosensitizers in the photoinactivation of microorganisms.

Funding: This work received financial support from PT national funds (FCT/MCTES, Fundação para a Ciência e Tecnologia and Ministério da Ciência, Tecnologia e Ensino Superior) through the projects LAQV-REQUIMTE (UIDB/50006/2020 and UIDP/50006/2020) and CESAM (UIDP/50017/2020+UIDB/50017/2020), and the FCT project POR2PS (FCT-EXPL/QUI-QOR/0586/2021). C Vieira thanks FCT for their doctoral grant (SFRH/PD/150358/2019). NMM Moura thanks FCT (Fundação para a Ciência e Tecnologia) for funding through program DL 57/2016—Norma transitória (Ref. CDL-CTTRI-048-88-ARH/2018).

Acknowledgments: The work was supported by University of Aveiro and FCT/MCTES through the projects UIDB/50006/2020 and UIDP/50006/2020 (LAQV-REQUIMTE) and (UIDP/50017/2020 and UIDB/50017/2020 (CESAM), funded by FCT/MCTES through national funds, and to the Portuguese NMR Network.

Conflicts of Interest: The authors declare no conflict of interest.

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