

Proceeding Paper

# Extraction and Complexing Abilities of a Series of p-Tert-butylcalix[4]arenes Substituted by N-carboxylmonoaza-12-crown-4 †

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**Abstract:** The extraction and complexing properties of a series of p-tert-butylcalix[4]arene derivatives containing N-methoxycarbonylmonoaza-12-crown-4 as substituents were evaluated through liquid-liquid extraction and spectrophotometric titration methods. It has been demonstrated that monosubstituted calixarene is an effective and highly selective extractant towards sodium cations. Both tri- and tetrasubstituted calixarenes form mononuclear complexes with sodium cations and complexes of 1:1 and 1:2 compositions (L-M) with barium and strontium cations. The ability of calix[4]arene with four crown ether substituents to form binuclear complexes with cations of transition metals - copper, nickel, cobalt and cadmium - has also been demonstrated.

**Keywords:** calixarene; crown ether; supramolecular architecture, complexing abilities, extraction

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## 1. Introduction

Calixarene derivatives are a new generation of highly selective molecular receptors capable of recognizing and binding ions and organic molecules of various natures. The introduction of a second complexing center, such as crown ether, into the calixarene molecule can make it possible to obtain compounds that demonstrate interesting steric and cooperative effects, which, in turn, can lead to increased complexing abilities and higher ion selectivity of the resulting supramolecular ensembles [1,2]. It should be noted that in the literature devoted to the properties of calixarenes, there is still a limited number of publications with systematic data on the dependence of the complexing properties of calixarenes on the nature and size of the intramolecular cavity formed by a different number of identical substituents or substituents representing fragments of molecules that make up homologous series.

Previously, we presented a comparison of the extraction properties of a series of p-tert-butylcalix[4]arene derivatives with crown-5 ether substituents – aza-15, benzoaza-15, amidobenzo-15-crown-5 [3]. This work presents a study of the extraction and complexing properties of the p-tert-butylcalix[4]arene series with N-carboxylmonoaza-12-crown-4 substituents at the lower rim. This choice is based on results demonstrating the ability of small crown ethers like 12-crown-4 to participate in the binding of not only s-elements cations but transition metals [4] and the well-known contribution of the amide group, which is part of the molecules of calixarene derivatives, to increasing stability complexes with metal salts [5,6].

## 2. Materials and Methods

### 2.1. Chemistry

Calixarenes **5**, **6** and **8** were synthesized according to the procedures described in [7,8].

The  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra were recorded from 10% solutions in chloroform-*d* on a Bruker Advance DRX 500 spectrometer using tetramethylsilane as internal reference. The mass spectra were recorded on an Agilent 6530 Accurate Mass Q-TOF spectrometer with LC/MS System. Absorption spectra in the UV region were recorded using a spectrophotometer SPECORD 250 Plus. All of metal chlorides were of analytical grade.

*5,11,17,23-tetra-tert-butyl-25,26,27-tris[(1,4,7-trioxa-10-aza-cyclododec-10-yl)carbonylmethoxy]-28-hydroxycalix[4]arene* (**7**). White solid; yield 85%.  $^1\text{H}$  NMR:  $\delta$  7.7 (s, 1H, OH), 7.2 (d,  $J = 2.56$  Hz, 2H, arom.), 7.12 (s, 2H, arom.), 7.08 (d, 2H, arom.), 6.79 (s, 2H, arom.), 4.89 (s, 6H,  $\text{CH}_2\text{CO}$ ), 4.52 (d,  $J = 13.45$  Hz, 2H,  $\text{ArCH}_2\text{Ar}$ ), 4.39 (d,  $J = 12.76$  Hz, 2H,  $\text{ArCH}_2\text{Ar}$ ), 4.0 (t, 6H,  $-\text{CH}_2\text{CH}_2\text{-N}$ ), 3.8 (t, 6H,  $-\text{CH}_2\text{CH}_2\text{-N}$ ), 3.65 (t, 6H,  $-\text{CH}_2\text{CH}_2\text{-N}$ ), 3.52 (m, 18H,  $\text{O-CH}_2\text{CH}_2\text{-O}$ ,  $-\text{CH}_2\text{CH}_2\text{-N}$ ), 3.23 (m, 4H,  $\text{ArCH}_2\text{Ar}$ ), 1.18 (s, 9H,  $(\text{CH}_3)_3\text{C}$ ), 1.14 (s, 18H,  $(\text{CH}_3)_3\text{C}$ ), 0.97 (s, 9H,  $(\text{CH}_3)_3\text{C}$ ).  $^{13}\text{C}$  NMR:  $\delta$  171.23 (C=O), 168.54 (C=O), 153.81, 153.11, 152.53, 146.82, 146.75, 142.48, 134.18, 134.10, 131.87, 126.89, 126.16, 72.24, 70.65, 70.43, 70.32, 64.78, 49.69, 44.62, 34.45, 34.79, 33.28, 31.86, 31.82, 31.76, 31.56, 29.87. MS ESI(+):  $m/z$  1293.7615  $[\text{M}+1]^+$ , 1316.7487  $[\text{M}+\text{Na}]^+$ .

### 2.2. Metal Picrate Extraction

After mutual saturation of the solvents, equal volumes (5 mL) of aqueous solutions of metal picrates ( $2.5 \times 10^{-4}$  M) and solutions of the calixarenes ( $2.5 \times 10^{-4}$  M) in  $\text{CHCl}_3$  were mechanically stirred for 1 h at 20 °C. After complete phase separation, the absorbance *A* of picrate ion remaining in the aqueous phase after extraction was determined spectrophotometrically at 350–380 nm. For each cation–calixarene system, the absorbance measurements were repeated at least four times. Blank experiments without calixarene were run under the same conditions, yielding an absorbance  $A_0$ . The percentages of extraction of metal picrates (%E) were calculated as the ratio  $100 \times (A_0 - A)/A_0$  as already described in details [9,10]. The results of preliminary measurements of the distribution coefficient *D* for calixarenes **5–8** in the concentration range  $4 \times 10^{-5}$ – $4 \times 10^{-4}$  M indicate the absence of “significant” solubility of calixarenes in water.

### 2.3. Stability Constant Determination

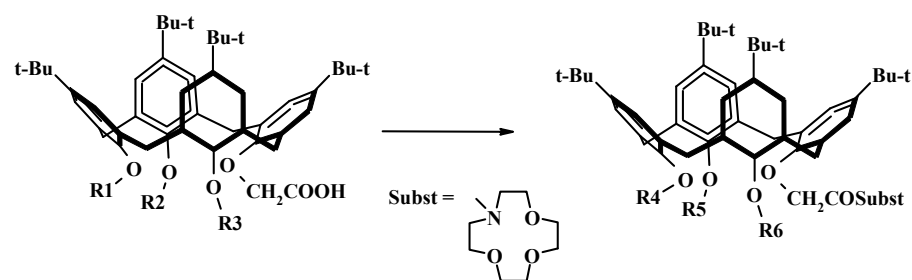
**UV–vis titration experiments.** A solution of calixarenes **5–8** (concentration about  $2 \times 10^{-5}$ – $3.5 \times 10^{-5}$  M) in methanol was treated with increasing amounts of metal chloride solution (concentration about  $1$ – $4 \times 10^{-4}$  M) containing proper ligand of the same concentration at 20 °C. The host concentration was maintained constant and the molar ratio of guest increased with respect to the host over the range 0.1:1 to 10(20): 1 during the titration. The absorbance measurements were carried out at six–ten wavelengths, at which spectral changes were the most notable (220–320 nm) simultaneously, and sets of the obtained experimental values ( $4 \times 21$  points) were used for joint computer processing [11,12]. The data were processed with the nonlinear least squares fitting SIRKO software [13].

In the case of studies involving transition metal salts, a solution of a ligand of a certain concentration was prepared with the addition of 0.01 M tris(hydroxymethyl)aminomethane (TRIS) in methanol.

## 3. Results and Discussion

To introduce substituents containing an amide fragment into the calixarene molecule, a previously developed approach based on the acylation of amino components with carboxy-substituted calixarenes (**1–4**) under the conditions of the activated ester method in the presence of DCC and hydroxybenzotriazole (HOBt) was implemented [3,7]. Thus,

the series was synthesized: mono- (5), di- (6), tri- (7) and tetra- (8) [(1,4,7-trioxa-10-azacyclododec-10-yl)carbonylmethoxy]-p-tert-butylcalix[4]arenes (Scheme 1).  $^1\text{H}$  NMR spectra indicate that molecules of all calixarenes adopt a cone conformation.



1: R1 = R2 = R3 = H; 2: R1 = CH<sub>2</sub>COOH, R2 = R3; 3: R1 = H, R2 = R3 = CH<sub>2</sub>COOH; 4: R1 = R2 = R3 = CH<sub>2</sub>COOH

5: R4 = R5 = R6 = H; 6: R4 = CH<sub>2</sub>COSubst, R5 = R6; 7: R4 = H, R5 = R6 = CH<sub>2</sub>COSubst; 8: R4 = R5 = R6 = CH<sub>2</sub>COSubst

**Scheme 1.** Synthesis of p-tert-butylcalix[4]arene derivatives 5-8 with N-carbonyl-monoaza-12-crown: DCC, HOBt, monoaza-12-crown-4, CHCl<sub>3</sub>, r.t., 12–25 h.

In order to evaluate the ability of calix[4]arenes with crown ether substituents to recognize metal ions, a liquid-liquid extraction experiment of some alkali, alkaline earth and transition metal picrates has been carried out. The obtained results are given in the Table 1 (Table 1).

**Table 1.** Extraction (%) of picrates by calixarene derivatives 5-8 at ligand-metal ratio-1:1.

Comp.	Li <sup>+</sup>	Na <sup>+</sup>	K <sup>+</sup>	Rb <sup>+</sup>	Cs <sup>+</sup>	Ca <sup>2+</sup>	Sr <sup>2+</sup>	Fe <sup>3+</sup>	Cu <sup>2+</sup>	Cd <sup>2+</sup>
5	- *	77.4	15.8	15.91	9.1	-	4.08	26.3	7.16	7.17
6	-	29.57	24.74	13.41	5.7	3.71	4.1	21.7	7.19	2.0
7	-	47.81	11.27	22.8	9.57	7.42	15	58.4	18.86	14.5
8	50.4	94.6	80.25	30.6	46.36	37.39	45.3	54.16	23.7	53.4

\* “-” –no extraction.

The results of studying the extraction properties of the calixarenes 5–8 indicate that monosubstituted calixarene 5 demonstrates high selectivity and efficiency toward to sodium cations, extracting sodium picrate by 77% ( $S_{\text{Na/K}}$  4.88) from aqueous solutions. And only with sodium cations is the process of complex formation with the participation of a monosubstituted macrocycle observed under spectrophotometric titration conditions. For di- 6 and tri- 7 substituted calixarenes, comparable and relatively low percent extraction values for most of the metals studied are observed. The exception is the results of the extraction of sodium and iron cations with trisubstituted calixarene 7, which are quite difficult to explain. Tetrasubstituted calixarene 8 is highly effective as an extractant of alkali and alkaline earth metals, as well as cadmium and iron (III) cations in the series of transition metals, however, this compound does not exhibit the properties of a selective extractant. This behavior can be explained by the large size of the intramolecular cavity formed by substituents containing donor centers of different nature, which can take part in the coordination of both “hard” and “soft” cations, and by the absence of serious steric hindrances in the interactions under consideration.

Note that in the course of studying the extraction properties toward to magnesium and barium cations for all four calixarenes, a re-extraction process is observed with a partial transition of the resulting complexes into the aqueous phase.

The results of studying the complexing properties of calixarenes toward s-element cations indicate the ability of disubstituted calixarene 6 to form 1:1 complexes with most of the metals studied. On the contrary, cesium cation forms with the disubstituted calixarene a biligand mononuclear complex with a high  $\lg K_{21} > 6$  value (Table 2).

**Table 2.** Stability constants ( $\lg K_n$ ) of the complexes of calixarenes 5-8 with alkali and alkaline earth metal cations in MeOH.

Comp.		Cation							
		Li <sup>+</sup>	Na <sup>+</sup>	K <sup>+</sup>	Rb <sup>+</sup>	Cs <sup>+</sup>	Ca <sup>2+</sup>	Sr <sup>2+</sup>	Ba <sup>2+</sup>
5	$\lg K_1$	- *	3.69	-	-	-	-	-	-
6	$\lg K_{11}$	-	2.31	2.98	2.57	2.27	2.62	2.98	3.66
	$\lg K_{21}$					6.67			
7	$\lg K_{11}$	-	>7	3.64	3.09	2.15	>7	5.15	5.24
	$\lg K_{12}$		-	2.73	2.57	-	-	2.67	3.27
8	$\lg K_{11}$	1.21	5.57	3.40	-	-	5.59	5.12	4.80
	$\lg K_{12}$	3.46						1.69	2.28

\* The changes in spectrum were too small and it was impossible to calculate the binding constant.

For trisubstituted calixarene 7, the highest values of stability constants of complexes in the series of alkali metals are observed for the sodium cation, with very high sodium selectivity—3.5–5 orders of logarithmic units. Complexes of calixarene 7 with alkaline earth metal cations are characterized by higher stability constants compared to the data obtained for the series of alkali metals (Table 2). At the same time, calcium selectivity is observed in the series of alkaline earth metals, which are about 2 logarithmic units. With sodium, cesium and calcium cations, calixarene 7 forms complexes of only 1:1 composition; in other cases, along with 1:1 complexes, there are complexes of 1:2 composition (L:M). Tetrasubstituted analogue 8 exhibits the same behavior as trisubstituted calixarene toward to cations of alkaline earth metals, but with lower stability constants of the complexes. It should be noted that only tetrasubstituted calixarene 8 acts as an effective extractant for lithium, forming 1:2 (L:M) complexes with lithium cations, while not demonstrating affinity for cations with a larger ionic radius - rubidium and cesium.

As is known, the optimal ratio of the effective diameter of the macrocycle to the ionic radius of the metal cation, which is 0.7–0.9:1, plays an important role in the process of complexation of crown ethers with cations of s-elements [14,15]. In the case of the interaction of crown ethers with transition metal salts, this model is not always realized, due to the possible participation of water molecules from the outer sphere in coordination with macrocycle heteroatoms [15]. Therefore, crown ethers exhibit low affinity for transition metal cations. The introduction of several aza-12-crown-4 ether residues into the pre-organized structure of calix[4]arene in the cone conformation will, in our opinion, make it possible to create a polytopic receptor for transition metal ions. Since the calix[4]arene molecule is a relatively rigid framework, and the crown ethers are connected to it by a rather short linker, the formation of sandwich or pseudosandwich type complexes with metal cations of large ionic radius can be expected. To test this assumption, we studied the interaction of tetrasubstituted calixarene 8 with transition metal cations by spectrophotometric titration. As follows from Table 3, tetrasubstituted calixarene molecules are capable to forming binuclear complexes with cations of copper, nickel, cobalt and cadmium, and the greatest stability of the complexes in this series is observed in the case of interaction with cadmium ions (Table 3).

**Table 3.** Stability constants ( $\lg K_n$ ) of the complexes of calixarene 8 with cations of some transition metals (ionic rad.) in MeOH.

$\lg K_n$	Cation					
	Fe <sup>3+</sup>	Ni <sup>2+</sup>	Cu <sup>2+</sup>	Co <sup>2+</sup>	Mn <sup>2+</sup>	Cd <sup>2+</sup>
	(1.28)	(1.38)	(1.44)	(1.48)	(1.60)	(1.94)
$\lg K_{11}$	2.88	2.80	4.95	3.24	3.02	4.81
$\lg K_{12}$		2.86	2.02	2.54		4.12

#### 4. Conclusions

Thus, a change in the extraction and complexing properties of calix[4]arenes with azacrown ether substituents-potential complexones-has been demonstrated, depending on the number of crown ether substituents included in the calixarene molecule.

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