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Polarization by the effect of a small torsional change in the benzothiazole (A)-benzobisthiazole (B) oligomer A-B13-A

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Abstract:

We use a method for the calculation of the molecular dipole ($\bar{\mu}$) and quadrupole ($\bar{\Theta}$) moments and dipole-dipole ($\bar{\bar{\alpha}}$), dipole-quadrupole ($\bar{\bar{A}}$) and quadrupole-quadrupole ($\bar{\bar{C}}$) polarizabilities which we have successfully applied to benzothiazole (A)-benzobisthiazole (B) linear oligomer A-B13-A. Two model rotational isomers have been characterized: 1) the fully-planar (000) conformation; and 2) a rotational isomer with each unit rotated in the range $1-10$¡ in the same direction (+++). For isomer 000, $\bar{\mu}$ is smaller than for +++. The calculation of $\bar{\bar{\alpha}}$, $\bar{\bar{A}}$ and $\bar{\bar{C}}$ has been carried out by the interacting induced dipoles polarization model that calculates tensor effective anisotropic point polarizabilities (method of Applequist *et al.*). The values of $\bar{\bar{A}}$ are specially sensitive to $\bar{\mu}$ which varies under rotation. This fact explains the great values of $|A_{x,xx}|$ for the polar isomer +++. This rotational isomer has the maximum hydrophilic accessible surface area which can improve solubility in water. It is found that small torsional changes can enhance solubility by increasing the hydrophilic accessible surface area. However, the torsion of the oligomer varies the value of $\bar{\mu}$ and so modifies $\bar{\bar{\alpha}}$, $\bar{\bar{C}}$, and, specially, $\bar{\bar{A}}$. For conformer +++, polarization effects occur even when each unit is rotated only $\phi-1$¡. The observed effect is an increase in $\bar{\mu}$ and, hence, in $\bar{\bar{\alpha}}$, $\bar{\bar{C}}$, and, specially, $\bar{\bar{A}}$.

Keywords: multipole moments, polarization, polarizability, geometric descriptors, partition coefficient.

Introduction

An important goal of computational chemistry research is the design of molecules for specific applications. Factors that must be taken into account include shape, size, electronic properties and reactivity. For any physical and chemical properties, the molecular shape largely determines the final function, and this is the rationale for conformational studies. The search for materials with large values of polarizabilities and other optical properties has attracted much interest from the experimental and the theoretical points of view.[1-3] An important consequence of various chemical processes involving charge redistribution is the changes in intermolecular interactions of the molecules involved.[4-7]

Molecular polarization is essential for the description of both optical properties and intermolecular interactions.[8-13] In addition, the intramolecular force fields used in molecular mechanics are usually based on the assumption that the interaction between nonbonded atoms can be analyzed in the same way as intermolecular interactions.[14-16] A possible improvement of conventional force fields could come from the introduction of an induction term.[17] Modern theoretical chemistry, with the help of powerful computers, can attempt to become a substitute for difficult experimental determinations.[18]

Classical models such as the dipole interaction model and its improvements, which were investigated among others by Applequist,[19,20] are able to predict molecular polarizabilities with surprising accuracy. Because of their simplicity they can be conveniently implemented in molecular force fields to account for the nonadditivity effects in the intermolecular interactions.[14-16,21-23]

In the theory of higher-order polarizabilities, when the external electric field is nonuniform the hamiltonian $\hat{H}(E_a, E_{ab}, \ldots)$ and the energy $W(E_a, E_{ab}, \ldots)$ may be expressed as a power series in the electric field, E_a, its gradient, E_{ab}, its hessian, E_{abc}, *etc.*:[24]

$$W(E_a, E_{ab}, \ldots) = W - \mu_a E_a - \frac{1}{3}\Theta_{ab}E_{ab} - \frac{1}{2}\alpha_{ab}E_a E_b - \frac{1}{3}A_{a,bc}E_a E_{bc} - \frac{1}{6}C_{ab,cd}E_{ab}E_{cd} - \ldots$$

where a, b, c,É denote Cartesian components, and the Einstein convention of summation over repeated indexes is implied here and in all subsequent equations. The electric multipole expansion is formed by the dipole ($\bar{\mu}$) and quadrupole ($\bar{\Theta}$) moments; the polarizability expansion is formed by the dipole-dipole (), dipole-quadrupole () and quadrupole-quadrupole () polarizabilities. In the following equations, a partial derivative with respect to E_{abc}É is the mathematical operation of taking the derivative with respect to E_{abc}É while keeping all other components of the field and its derivatives constant. The total electric multipole moments are

In previous articles[25-31] we proposed a method for the calculation of , , and that was applied for the study of benzothiazole (A)-benzobisthiazole (B) linear oligomers.[26-28] In Reference 26 we calculated fully-planar linear oligomers in the series A-B1-A to A-B13-A and extrapolated the results to give predictions for polybenzobisthiazole A-B�-A. This choice was prompted by the importance of this heteroaromatic structure in rodlike polymeric chains that possess interesting mechanical and optical properties.

It should be noted that polymers do not usually have well-defined molecular structures.[18] A sample of a polymer will contain molecules having possibly a variety of structures. Rational polymer design technologies rely on a detailed structural knowledge of oligomers at the atomic level. Although the strong C-C, C-N and C-S bonds maintain the integrity of the polymer, A-B�-A has a high degree of flexibility because rotation can occur about the intercycle bonds between monomeric units in the polymer chain. Thus the polymer chain can adopt a large number of configurations in three-dimensional space. The torsion angles of intercycle bonds play a crucial part in determining properties, for these determine the overall polarity of the macromolecule. In Reference 27 we examined the torsional effects on $\bar{\bar{\alpha}}$, $\bar{\bar{A}}$ and $\bar{\bar{C}}$ for A-B13-A. We characterized three model rotational isomers: (1) the fully-planar (000) conformation; (2) a rotational isomer where each unit is rotated 10¡, rotations being performed in the alternate directions (+-+); and (3) a rotational isomer where each unit is rotated 10¡, all rotations being performed in the same direction (+++). We dealt with small torsion angles because the molecular polarizabilities were due to p-p overlap between successive monomeric units in the one-dimensional chain and this overlap falled sharply for big rotations. The main conclusion of this previous work was that small torsional changes in A-B13-A affect $\bar{\bar{\alpha}}$ and $\bar{\bar{C}}$, and, specially, $\bar{\bar{A}}$. This is easy to understand because $\bar{\mu}$ is changed by rotation and $\bar{\bar{A}}$ is highly sensitive to $\bar{\mu}$. This previous work allowed us to ask ourselves whether changes in $\bar{\mu}$, $\bar{\bar{\alpha}}$, $\bar{\bar{A}}$ and $\bar{\bar{C}}$ holds for smaller angle rotations. So, in this article we have studied the polarizabilities of the fully-planar 000 and the +++ oligomers where each unit is rotated from 1¡ to 10¡.

Electrostatic properties

Atomic net charges and polarizabilities are calculated from their s and p contributions. The s net charges and polarizabilities are estimated by the principle of electronegativity equalization[32-35] but applied bond by bond in the molecule.[25-31] The p net charges and polarizabilities are evaluated with the method of HŸckel. Dipole ($\bar{\mu}$) and tensor quadrupole ($\bar{\Theta}$) moments have been calculated from the point (atomic) distribution of net charges. The molecule is brought into its principal inertial coordinate system.

It is well-known that intercycle p conjugation vanishes for perpendicular cycles. So, HŸckel b parameter should decrease for out-of-plane conformations, tending to zero for perpendicular polycyclic structures. Therefore, if the functional relation of p conjugation parameters with respect to the dihedral angles is known, improvements can be made in the simple HŸckel calculations.[36] Mulliken[37,38] has proposed that HŸckel b parameter be proportional to the overlap integral S and has shown that in diatomic systems the proportionality holds rather well at bond distances of interest. For Slater overlap integrals[39] b as a function of distance has been calculated.[40]

Although there is no convincing evidence as to which b-distance curve to use, Mulliken�s success with diatomic molecules[37,38,41] is suggestive, as is a successful application to calculation of ionization potentials.[42] Thus, in first approximation, b is proportional to the overlap integral S between two fragments. This is certainly a crude approximation but often used to deduce the tendency in the variation of b. It should be noted that S between two parallel orbitals with no spherical symmetry varies as the cosine of their relative torsion angle q. Acceptance of the assumption further implies that b can be evaluated, in first approximation, between p$_z$ orbitals twisted from coplanarity by an angle q as[43-46]

$$\beta = \beta_0 \cos\theta$$

Thus, it is convenient to define a standard b$_0$ (being equal to the b parameter for benzene) for the planar q dihedral angles and to modify other b parameters (in units of the standard b$_0$) with the use of this

dimensionless function.

Of course, this method is not exact. It should be remarked that the rotation of a torsion angle can generate conformational changes. For example, the bond distance can vary and the overlap between two atomic orbitals decreases exponentially when the distance between them increases. So, the overlap integral S not only depends on the relative orientation of the p_z orbitals directly involved in the fragment-fragment bonding but also on second-neighbour overlaps between p_z orbitals belonging each to a different fragment. The proper b function in this case is considerably different from the cosine function being used. An improved estimate of the b function can be expected from some kind of empirical information. Joachim et al.[47] have evaluated the electronic coupling, V_{ab}, of the binuclear mixed valence M_{II}-L-M_{III} complex $[(NH_3)_5Ru\text{-bipyridyl-}Ru(NH_3)_5]_{5+}$ in the valence bond electronic states M_{II}-L-M_{III} (a) and M_{III}-L-M_{II} (b). When a pyridine ring rotates around the ligand axis, p-p can be best fitted by a function. From this observation we assume that the b function is universal and has the same form as the electronic coupling function, V_{ab}, for complex $[(NH_3)_5Ru\text{-bipyridyl-}Ru(NH_3)_5]_{5+}$:

Interacting induced dipoles polarization model for molecular polarizabilities

The computation of molecular polarizabilities has been carried out by the interacting induced dipoles polarization model presented elsewhere (14-16,25-28). It calculates tensor effective anisotropic point polarizabilities by the method of Applequist *et al*.[19,20,48-50] One considers the molecule as being made up of N atoms *(i, j, k,É)*, each of which acting as a point particle located at the nucleus and responding to an electric field only by the induction of a dipole moment which is a linear function of the local field. If a Cartesian component of the field owing to the permanent multipole moments is , then the induced moment in atom *i* is

(1)

where a_i is the polarizability of atom *i* and is the symmetrical field gradient tensor, , where *e* is the charge of the proton and the subscripts *a*, *b*, *c*,É stand for the Cartesian components *x*, *y*, *z*. The set of coupled linear Equations (1) for the induced dipole moments can conveniently be expressed in compact matrix equation form, if one introduces the matrices and , with elements and (being the Kronecker d*)*, respectively. To suppress the restriction in the sum, the diagonal elements are defined as zero. Similarly, and are column vectors with elements and . Equation (1) is thus written,

where is the -dimensional unit matrix. This matrix equation can be solved for the induced dipoles as

Here the symmetrical many-body polarizability matrix, , has been introduced:

The compact matrix equation $\bar{\mu} = \bar{\bar{A}}\bar{E}$ is equivalent to the N matrix equations:

$$\bar{\mu}^i = \sum_{j=1}^{N} \bar{\bar{A}}^{ij}\ \bar{E}^j$$

Let the molecule be in a uniform applied field, so that $\bar{E}^j = \bar{E}$ for all *j*. Then

$$\bar{\mu}^i = \left[\sum_{j=1}^{N}\bar{\bar{A}}^{ij}\right]\bar{E} = \bar{\bar{\alpha}}^{eff,i}\ \bar{E}$$

The coefficient of $\bar{E}$ in this equation is seen to be an effective (hence, additive) polarizability of unit *i*, $\bar{\bar{\alpha}}^{eff,i}$. The total moment induced in the molecule $\bar{\mu}^{mol}$ is

$$\bar{\mu}^{mol} = \sum_{i=1}^{N}\bar{\mu}^i = \left[\sum_{i=1}^{N}\sum_{j=1}^{N}\bar{\bar{A}}^{ij}\right]\bar{E} = \left[\sum_{i=1}^{N}\bar{\bar{\alpha}}^{eff,i}\right]\bar{E}$$

from which it is seen that the molecular polarizability tensor $\bar{\bar{\alpha}}^{mol}$ is

$$\bar{\bar{\alpha}}^{mol} = \sum_{i=1}^{N}\sum_{j=1}^{N}\bar{\bar{A}}^{ij} = \sum_{i=1}^{N}\bar{\bar{\alpha}}^{eff,i} \qquad (2)$$

The most noticeable feature of molecular polarizability surfaces is a curve of discontinuity along which the polarizability approaches �� under certain conditions.[51,52] The significance of a polarizability of �� is that the molecule is in a state of resonance and absorbs energy from the applied field. This behaviour of the model can be understood from its close relation to the classical system of N coupled oscillators, which likewise shows resonance under conditions other than the resonance conditions of the isolated oscillators.[53-56]

Voisin[57] wrote a computer program called PAPID for the computation of molecular polarizabilities. A data base of atomic polarizabilities has been built up from coupled Hartree-Fock calculations of a series of molecules containing the main functional groups which are presented in organic molecules,[17] and has been implemented in the Block Data of the program.[58,59] In this work, we have implemented the following improvements in PAPID.

1. We have used a damping function in the calculation of the symmetrical field gradient tensor in order to prevent the polarizability from going to infinity.[17] Thus, if the distance *r* between atoms *i* and *j* is lesser than the cut-off value, $s = 1.662\left(\alpha^i\alpha^j\right)^{1/6}$, the field gradient tensor is now evaluated as

$$\bar{\bar{T}}^{ij} = -\frac{3v^4}{r^5}\begin{bmatrix} x^2 - \frac{4v^3-3v^4}{3v^4}r^2 & xy & xz \\ xy & y^2 - \frac{4v^3-3v^4}{3v^4}r^2 & yz \\ xz & yz & z^2 - \frac{4v^3-3v^4}{3v^4}r^2 \end{bmatrix}$$

where $v = r/s$.

2. We have neglected the interaction between bonded atoms and between whichever atoms with a distance lying out of an interval defined by .[17] The starting values for this interval are and we increment r_{inf} whether reasonance conditions are detected. Furthermore, we have improved this feature using the following algorithm:[25-28]

Step 1: Set r_{inf} to zero.

Step 2: Neglect interaction for atoms with distance lying in the interval .

Step 3: If a test indicates resonance conditions, then increment r_{inf} by 1 a.u. and continue from *Step 2*.

Step 4: Stop.

After the application the algorithm, the typical value of r_{inf} is 6 a.u. Since the results depend on the choice of these parameters, which in fact control the interaction between atoms, we have minimized this effect using a value for r_{sup} which is never reached, and a value for r_{inf} which is the minimum allowed for each model and molecule, compatible with the condition of no resonance.

3. We have also implemented the following tests indicating a resonance condition:[25-28] (a) test whether the matrix is singular:[17] ; (b) test whether the matrix is not defined positive *(i.e*., some eigenvalue is less than or equal to zero);[17] (c) test whether the matrix is not defined positive; and (d) test whether some effective matrix is not defined positive.

The values of are calculated with this model. We give particular attention to the average molecular polarizability and the anisotropy g.[17]

A fully operative version of the PAPID program[57] including the whole interacting induced dipoles polarization model has been implemented[30,31] into the MM2 program for molecular mechanics.[60,61] We have called the new version MMID.[14-16] A description of the implementation has been reported◆ elsewhere.[25]

Geometric descriptors and topological indices

The TOPO program for the theoretical simulation of molecular shape has been described elsewhere.[25-27,30,31,62] The surface of molecules can be represented by the external surface of a set of overlapping spheres with appropriate radii, centred on the nuclei of the atoms.[63-67] If the radii are those of van der Waals,[68] the *bare* (solvent free) molecular surface is obtained;[63,64] if the radii are those of van der Waals plus the effective radius of a solvent molecule, the solvent-accessible molecular surface is now obtained.[65]

The characterization of the molecular shape is helped by the calculation of geometric descriptors and topological indices of the molecules.[63,64] The molecule is treated as a solid in space defined by tracing spheres about the atomic nuclei. It is computationally enclosed in a graduated rectangular box and the geometric descriptors evaluated by counting points within the solid or close to chosen surfaces. The molecular volume V is concurrently approximated as , where P is the number of points within the molecular volume (within distance R_x of any atomic nucleus X) and GRID is the size of the mesh grid.

As a first approximation, the *bare* molecular surface area could be calculated as , where Q is the number of points close to the *bare* surface area (within a distance between Rx and of any atomic nucleus X). However, the estimate has been improved:[64] if the point falls exactly on the surface of one of the atomic spheres, it accounts indeed for $GRID^2$ units of area on the *bare* molecular surface. This is because the total surface of atom X can accommodate points. When a point falls beyond the surface, it represents $GRID^2$ units of area on the surface of a sphere of radius , not on the surface of atom X. On the surface of X it accounts only for a fraction of this quantity, namely, . The total bare surface area is, therefore, calculated as , where F is the sum of elements AF defined as for those points close enough to the surface of any atom X. is the squared radius of atom X and is the squared distance of point I from the atomic nucleus X. Meyer[64] has written a program including the OEPP (one element per point) subprogram that allows the calculation of the molecular volume and bare surface area. We have written a version of this subprogram (we have called it ATOEPP) that allows the atom-to-atom partition analysis of these geometric descriptors into their atomic component parts.[25-31]

Two topological indices of molecular shape can be now calculated: G and G◆.[64] The ratio has been interpreted as a descriptor of molecular globularity; S_e is the surface area of a sphere whose volume is equal to the molecular volume V. The ratio has been interpreted as a descriptor of molecular rugosity.

The importance of the solvent validates the assumption that the properties of the systems solvated in water are strongly related to the contact surface between solute and water molecules.[65,67] We can propose a new molecular geometric descriptor that is the *solvent-accessible surface* AS.[65] This surface is denoted when a spherical probe is allowed to roll on the outside while maintaining contact with the bare molecular surface. The continuous sheet defined by the locus of the centre of the probe is the accessible surface. Alternatively, the accessible surface is calculated as the bare molecular surface by using pseudo-atoms, whose van der Waals radii were increased by the radius R of the probe. The *accessibility* is a dimensionless quantity varying between 0 and 1 and represents the ratio of the accessible surface area in a particular structure to the accessible surface area of the same atom when isolated from the molecule.[65]

Fractal surfaces[69] provide a means for characterizing the irregularity of molecular surfaces.[70] The area of the solvent-accessible surface AS depends on the value of the probe radius R. The fractal dimension D of the molecules may be obtained according to Lewis and Rees[70] as

The fractal dimension, D, provides a quantitative indication of the degree of surface accessibility toward different solvents. The larger the D value the faster the accessible surface area drops with an increase in solvent molecular size.

A version of the TOPO program has been implemented into the AMYR program for the theoretical simulation of molecular associations and chemical reactions.[8-12] The algorithm allows for the characterization of molecular units and aggregates. All of the geometric descriptors and topological indices but fractal dimension can be now calculated.

Calculation results and discussion

Table 1. Electrostatic properties for the benzothiazole (A)-benzobisthiazole (B) A-B[13]-A linear oligomer.

f[a]	m[b]	Q[c]	Q1[d]	Q2	Q3	k[e]
0	1.456	-77	-481	0	250	-0.317
1	12.125	-77	-345	22	92	-0.681
2	15.190	-77	-587	21	334	-0.319
3	7.149	-77	-893	15	645	-0.181
4	6.441	-78	-757	10	514	-0.205
5	14.883	-78	-296	17	46	-0.832
6	12.448	-78	-705	21	452	-0.255
7	1.706	-77	-913	10	671	-0.166
8	11.533	-77	-587	16	340	-0.301
9	15.228	-77	-399	20	148	-0.532
10	7.760	-77	-871	15	626	-0.184

[a]Torsion angle (degrees) of each unit. All rotations are performed in the same direction.

[b]Dipole moment (debye).

[c]Mean quadrupole moment (D◆).

[d]Quadrupole moment tensor eigenvalues Q_1, Q_2 and Q_3 (D◆).

[e]Ray asymmetry parameter.

Two model rotational isomers of the benzothiazole (A)-benzobisthiazole (B) linear oligomer A-B[13]-A have been characterized: (1) the fully-planar (000) conformation; and (2) a rotational isomer in which each unit is rotated by a fixed angle f with respect to the previous one. All rotations are performed in the same direction (+++, see Figure). The fixed rotation angle f has been varied from 1¡ to 10¡. The dipole ($\vec{\mu}$) and quadrupole tensor ($\bar{\bar{\Theta}}$) moments for the rotational isomers are reported in Table 1. The geometries of the oligomers have been taken from the optimization of the monomeric units with MMID (molecular mechanics with interacting dipoles polarization), a molecular mechanics computer program[14-16] based on the MM2 parametrization.[60,61] All the linear oligomers are arranged in an all-*trans* configuration, in which the series of N=C double bonds, either in the same or in consecutive units, are disposed alternatively. Each rotational isomer is brought into its principal inertial coordinate system. The length x of the oligomer is defined as the maximal length (corresponding to the translation axis in A-B◆-A), the height z as its minimal thickness, and its width y is measured at right angles to the axes indicated by the length and height. The origin is the centre of mass for each conformation.

Figure. Image of fragment A-B-B- of the benzothiazole (A)-benzobisthiazole (B) linear oligomer A-B13-A: a) front view of the 000 rotational isomer; b) top view of the 000 rotational isomer; c) top view of the +++ rotational isomer. For clarity, the rotation angles have been slightly enhanced (to 20¡) in this figure to make clearer the torsional effects.

Two important features in Table 1 must be remarked. First, the sharp raising of the calculated dipole moments m on going from the planar (000) to the smallest distortion (û), as well as their oscillatory behaviour as f increases monotonically. Second, the values of the mean quadrupole tensors *(viz.,* the trace of) remain roughly constant about -77 D◆ in spite of the great variation of the three diagonal components. In other words, the anisotropy of changes as f increases, but a macroscopic measure of Q in a fluid phase would not detect such a change. As a measure of the anisotropy of we have calculated a Ray◆s type asymmetry parameter k71 in the last column of Table 1. k is defined from as:

which becomes -1 for a *prolate* symmetric quadrupole tensor () and 1 for an *oblate* symmetric quadrupole (), varying between these two values for asymmetric cases. In all of the cases in Table 1 , indicating a closer similarity of the molecular quadrupoles with a prolate symmetric tensor.

Table 2. Molecular dipole-dipole polarizabilities for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer.

fa	ab	a1c	a2	a3	axxd	axy	axz	ayy	ayz	azz
0	369	134	280	693	693	3	0	280	0	134
1	369	172	242	693	693	3	0	242	0	172
2	764	242	321	1728	1728	7	10	321	0	242
3	762	259	306	1720	1720	2	6	306	-3	259
4	759	253	314	1711	1711	3	3	314	1	253
5	758	260	308	1705	1705	4	5	308	-5	260
6	758	262	307	1705	1705	2	5	307	1	262
7	760	258	311	1711	1711	1	3	311	-4	258
8	766	267	302	1727	1727	2	3	302	-0	267
9	775	258	311	1755	1755	2	4	311	-3	258
10	788	265	303	1797	1797	0	4	303	-3	266

aTorsion angle (degrees) of each unit. All rotations are performed in the same direction.

bMean molecular dipole-dipole polarizability (Œngstršm3).

cMolecular dipole-dipole polarizability tensor eigenvalues a1, a2 and a3 ().

dMolecular dipole-dipole polarizability tensor Cartesian components (3).

The molecular , and tensors for the rotational isomers of A-B13-A have been calculated with PAPID program17 and the results are summarized in Tables 2 to 4. In general, the values of increase with the torsion angle f. The eigenvector analysis shows that this augmentation is basically due to the third eigenvalue a3 (see the fifth column in Table 2), which corresponds well to the axx Cartesian component, *i.e.*, the component along the chain axis. The results show that the calculated anisotropies g of the tensors increase, in general, with the torsion angle f (from 502 3 till 1513 3). The oligomers are reasonably predicted to be more polarizable along the chain direction than across the bonds (). The values of the higher order tensors and , reported in Tables 3 and 4, are also anisotropic, also showing more polarizability along the chain direction (, and).

Table 3. Molecular dipole-quadrupole polarizabilities for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer.

fa	Axb	Ay	Az	Ax,xxc	Ax,xy	Ax,xz	Ax,yy	Ax,yz	Ax,zz	Ay,xx	Ay,xy
0	-75	-19	0	-163	-55	0	81	0	81	-63	44
1	70	10	15	147	27	60	-90	17	-57	31	-1359
2	581	-77	62	844	-367	262	-426	15	-418	-486	-1728
3	379	-92	-30	492	-421	-95	-291	15	-201	-552	-2426
4	230	-19	-68	223	-142	-244	-125	30	-97	-212	-1458
5	224	5	-26	182	-54	-92	-121	-3	-60	-107	-1857
6	449	-30	-22	504	-180	-90	-281	34	-222	-251	-1701
7	988	-23	-64	1317	-129	-239	-671	1	-645	-178	-1276
8	1928	18	-58	2755	11	-195	-1425	19	-1331	-19	-2021
9	3361	8	-27	4953	-50	-97	-2475	18	-2478	-103	-1002
10	5350	-10	-43	8003	-82	-189	-4051	-15	-3951	-119	-1807

Table 3 (…/…)

Table 3 (…/…)

fa	Ay,xz	Ay,yy	Ay,yz	Ay,zz	Az,xx	Az,xy	Az,xz	Az,yy	Az,yz	Az,zz
0	0	28	0	35	0	0	6	0	-1	0
1	2532	-13	-1	-18	74	2537	1317	-38	1	-36
2	183	248	-3	238	333	198	1756	-168	3	-165
3	885	281	-4	272	-105	902	2503	51	3	55
4	776	111	-3	101	-290	792	1580	144	2	146
5	-55	59	-3	47	-113	-40	2012	57	3	57
6	899	131	-3	120	-119	912	1871	57	3	61
7	56	94	-4	84	-291	67	1441	143	1	148
8	322	13	-4	5	-225	330	2159	113	3	113
9	677	57	-2	45	-119	682	1091	61	4	58
10	-631	69	-2	50	-247	-628	1830	121	-2	126

aTorsion angle (degree) of each unit. All rotations are performed in the same direction.

bMolecular dipole-quadrupole polarizability x vector component (4).

cMolecular dipole-quadrupole polarizability tensor Cartesian components (4).

It can be seen from Table 4 that the dominant components of increase, in general, with the torsion angle f. Instead, the behaviour of is rather complicated as exemplified by the $A_{x,xx}$ component. While and are sensitive to the length of the oligomer, is expected to be also sensitive to the molecular dipole moment m, but no clear correlation can be estimated between the behaviour of the calculated m and components of .25-28

Table 4. Molecular quadrupole-quadrupole polarizabilities for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer.

fa	Cxx,xxb	Cxx,xy	Cxx,xz	Cxx,yy	Cxx,yz	Cxx,zz	Cxy,xy	Cxy,xz	Cxy,yy	Cxy,yz
0	1038000	2000	0	-519000	0	-518000	269000	0	-1000	0
1	1037000	2000	1000	-519000	0	-518000	208000	-2000	-1000	0
2	2344000	-4000	7000	-1171000	0	-1173000	338000	-31000	2000	0
3	2334000	-5000	-1000	-1168000	-1000	-1166000	297000	-26000	2000	0
4	2327000	3000	-3000	-1163000	0	-1164000	327000	-11000	-2000	0
5	2325000	4000	4000	-1162000	-1000	-1162000	313000	-27000	-2000	0
6	2332000	-1000	6000	-1166000	0	-1166000	312000	-5000	0	0
7	2352000	-1000	1000	-1175000	-1000	-1177000	324000	-22000	0	0
8	2389000	3000	2000	-1195000	0	-1194000	300000	-10000	-2000	0
9	2445000	1000	7000	-1221000	-1000	-1224000	325000	-14000	-1000	0
10	2523000	-3000	5000	-1262000	-1100	-1261000	305000	-20000	1000	0

Table 4 (…/…)

fa	Cxy,zz	Cxz,xz	Cxz,yy	Cxz,yz	Cxz,zz	Cyy,yy	Cyy,yz	Cyy,zz	Cyz,yz	Cyz,zz	Czz,zz
0	-1000	127000	0	0	0	260000	0	259000	0	0	259000
1	-1000	188000	-1000	0	-1000	260000	0	259000	0	0	259000
2	2000	206000	-4000	0	-4000	586000	0	586000	0	0	587000
3	3000	247000	1000	0	1000	585000	0	583000	0	0	583000
4	-1000	219000	2000	0	1000	582000	0	581000	0	0	583000
5	-2000	234000	-2000	0	-2000	581000	1000	581000	0	1000	582000
6	1000	235000	-3000	0	-3000	583000	0	583000	0	0	584000
7	1000	22300	-1000	0	-1000	588000	1000	588000	0	1000	589000
8	-1000	246000	-1000	0	-1000	598000	0	597000	0	0	597000
9	-1000	221000	-3000	0	-3000	611000	0	611000	0	0	613000

| 10 | 2000 | 240000 | -2000 | 0 | -2000 | 632000 | 1000 | 630000 | 0 | 1000 | 631000 |

[a]Torsion angle (degree) of each unit. All rotations are performed in the same direction.

[b]Molecular quadrupole-quadrupole polarizability tensor Cartesian components (5).

The calculation of the geometric descriptors and topological indices for A-B13-A in Tables 5 and 6 is compared with reference calculations carried out with the GEPOL (GEometry of POLyhedron) program.[72-75] This program has been developed in this laboratory by Silla *et al*. The algorithm calculates three kinds of envelope surfaces: (1) the *van der Waals molecular surface*, which is the external surface resulting from a set of spheres centred on the atoms or groups forming the molecule; (2) the *solvent-accessible surface*, which is the surface generated by the centre of the solvent, considered as a rigid sphere, when it rolls around the van der Waals surface;[65] and (3) the *solvent-excluding surface* which is composed of two parts: the contact surface and the reentrant surface. The *contact surface* is the part of the van der Waals surface of each atom that is accessible to a probe sphere of a given radius. The *reentrant surface* is defined as the inward-facing part of the probe sphere when this is simultaneously in contact with more than one atom. The GEPOL program is fast and efficient.

The geometric descriptors for the rotational isomers of A-B13-A are listed in Table 5. The value of the molecular volumes V remain roughly constant with the torsion angle f. The values compare well with reference calculations carried out with the GEPOL program. However, the molecular surface areas S slightly increase with f. The water-accessible surface areas AS, the hydrophilic HLAS contribution, and the side-chain accessible-surface areas AS♦, rather increase with f. However, the hydrophobic HBAS component part of AS rather decreases with f. It should be noted that large values of HLAS (and low values of HBAS), *e.g.* for ¡, favour solubility in water, which is fundamental for the processibility of the material.

Table 5. Geometrical descriptors for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer.

f[a]	V (3)[b]	V ref.[c]	S (2)[d]	S ref.[c]	AS (2)[e]	AS ref.[c]	HBAS[f]	HLAS[g]	AS' (2)[h]	AS' ref.[c]
0	1961.6	2065.5	1990.29	2175.68	3032.06	3190.15	2534.48	497.58	4990.88	5176.78
1	1963.0	2062.7	1993.10	2174.33	3023.22	3195.07	2523.50	499.72	4993.93	5193.09
2	1963.5	2065.8	1991.09	2175.62	3029.20	3199.66	2523.99	505.21	4999.17	5201.85
3	1963.4	2065.8	1995.45	2177.65	3034.91	3206.01	2521.72	513.19	5005.15	5209.04
4	1964.0	2064.6	1994.90	2179.67	3042.71	3210.47	2518.41	524.30	5018.99	5220.71
5	1963.6	2065.1	1998.11	2181.10	3045.54	3217.17	2510.94	534.60	5033.82	5230.11
6	1963.8	2065.2	1997.14	2181.72	3050.97	3222.29	2504.34	546.63	5040.88	5236.36
7	1964.0	2064.5	1998.89	2180.92	3060.90	3227.88	2501.17	559.73	5046.71	5243.50
8	1963.6	2063.8	1997.89	2179.85	3057.67	3232.42	2489.55	568.12	5051.15	5246.59
9	1962.9	2066.7	1998.95	2178.69	3062.50	3229.62	2487.22	575.28	5049.68	5247.45
10	1962.7	2062.5	1994.11	2177.48	3055.78	3227.36	2475.35	580.44	5047.17	5244.73

[a]Torsion angle (degree) of each unit. All rotations are performed in the same direction.

[b]Molecular volume (3).

[c]Reference: calculations carried out with the GEPOL program.

[d]Molecular surface area (2).

[e]Water accessible surface area (2).

[f]Hydrophobic accessible surface area (2).

[g]Hydrophilic accessible surface area (2).

[h]Side-chain accessible-surface area (2).

The topological indices for the rotational isomers of A-B13-A are reported in Table 6. The globularity index G is roughly constant with the torsion angle f. The values compare well with reference calculations carried out with the GEPOL program. The behaviour of the rugosity index G♦, the fractal dimension of the solvent-accessible surface D, and the fractal dimension averaged for *nonburied* atoms (those with nonzero accessible surface area) D♦, is the same as has been noticed for G. The comparison between the TOPO and GEPOL programs is of special interest because the latter does not allow an atom-to-atom partition analysis of the geometric descriptors and topological indices.[25-31]

Table 6. Topological indices for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer.

f[a]	G[b]	G ref.[c]	G'[d]	G' ref.[c]	D[e]	D ref.[c]	D'[f]
0	0.381	0.360	1.015	1.053	1.52	1.53	1.56
1	0.380	0.360	1.015	1.054	1.51	1.53	1.55
2	0.381	0.361	1.014	1.053	1.51	1.53	1.56
3	0.380	0.360	1.016	1.054	1.51	1.53	1.56
4	0.380	0.360	1.016	1.056	1.51	1.53	1.56
5	0.380	0.360	1.018	1.056	1.51	1.53	1.56
6	0.380	0.359	1.017	1.056	1.51	1.53	1.56
7	0.379	0.360	1.018	1.056	1.51	1.53	1.57

8	0.380	0.360	1.017	1.056	1.51	1.53	1.57
9	0.379	0.360	1.018	1.054	1.52	1.53	1.57
10	0.380	0.360	1.016	1.056	1.51	1.53	1.57

ₐTorsion angle (degree) of each unit. All rotations are performed in the same direction.

ᵦMolecular globularity.

ᵧReference: calculations carried out with the GEPOL program.

ₑMolecular rugosity ($^{-1}$).

ₑFractal dimension of the solvent-accessible surface.

ₑFractal dimension of the solvent-accessible surface averaged for non-buried atoms.

The atom-to-atom partition for each thiazolic nitrogen atom in the central unit of the oligomer are reported in Table 7. The contribution of this nitrogen to the molecular volume V slightly increases with the torsion angle f. This torsional effect has a steric cause and is due to the partial opening of the oligomer structure underwent under rotation. The trend of the contribution of the N atom to the molecular surface areas S, the molecular rugosities $G\blacklozenge$ and the water-accessible surface areas AS, follows that for the molecular volume. Thus, the *accessibilities* of the N atom surfaces increase with f (till 21.2%). It should be noted that the polar thiazolic nitrogens are the only atoms that contribute to the hydrophilic accessible surface area HLAS, and consequently, the conformer with ¡ shows the greatest value of HLAS in Table 5. The contribution of the N atoms to the molecular globularities G and to the fractal dimensions $D\blacklozenge$, decrease, in general, with f. The N atoms contributions to $D\blacklozenge$ are greater () than for the molecular values averaged for nonburied atoms (in the last column of Table 6). This means that the irregularity of the molecular surface near the N atom is greater than on average on the whole molecule.

Table 7. Geometrical descriptors and topological indices for the benzothiazole (A)-benzobisthiazole (B) A-B13-A linear oligomer. Atom-to-atom partition for each thiazolic nitrogen atom in the central unit.

fₐ	V (^{3})ᵦ	S (^{2})ᵧ	Gₑ	G� ($^{-1}$)ₑ	AS (^{2})ᵩ	accessibilityₘ	D�ₕ
0	9.2	12.13	1.752	1.317	18.11	0.180	1.81
1	9.2	11.96	1.776	1.299	17.99	0.179	1.79
2	9.2	11.97	1.777	1.297	18.49	0.184	1.80
3	9.2	12.19	1.743	1.323	18.41	0.183	1.76
4	9.2	12.28	1.733	1.330	18.81	0.187	1.78
5	9.3	12.34	1.727	1.334	19.26	0.191	1.72
6	9.2	12.57	1.696	1.359	19.92	0.198	1.69
7	9.2	12.64	1.685	1.367	19.73	0.196	1.68
8	9.3	12.50	1.706	1.350	20.22	0.201	1.66
9	9.2	12.52	1.702	1.354	20.69	0.206	1.62
10	9.3	12.39	1.720	1.339	21.32	0.212	1.62

ₐTorsion angle (degree) of each unit. All rotations are performed in the same direction.

ᵦMolecular volume (^{3}).

ᵧMolecular surface area (^{2}).

ₑMolecular globularity.

ₑMolecular rugosity ($^{-1}$).

ᵩWater accessible surface area (^{2}).

ₘAccessibility of the accessible surface.

ₕFractal dimension of the solvent-accessible surface.

Conclusions

In this article we use a method for the calculation of molecular and moments and , and polarizabilities that we have successfully applied to two model rotational isomers of the A-B13-A linear oligomer.[14-17,25-28] In a previous work,[27] we studied the torsion for bigger angles (¡). This work allowed us to ask ourselves whether a small torsion f (a few degrees) would produce appreciable effects similar to those detected for bigger f. The objective of this article is to see whether a minimal distortion (¡) from the planar structure produces the same effect than a bigger distortion (¡). The main conclusion of this work is that, in fact, small distortions of the strictly planar molecular geometry lead to significant changes in some of the studied properties. In particular, specially sensitive properties are , , a, a$_{xx}$, A_x, $A_{x,xx}$, $C_{xx,xx}$, AS, HBAS, and HLAS. In turn, , , A_x, and $A_{x,xx}$ does not show a monotonic change with f. Remarkably, , , a, a$_{xx}$, A_x, $A_{x,xx}$, and $C_{xx,xx}$ are sensitive to a change in f of a few degrees. Small torsional changes in the A-B13-A linear oligomer can enhance solubility by increasing the hydrophilic accessible surface area. This change is specially related with the accessiblity of the N atoms.

Additional conclusions follow.

1. The presented model that is based in the electronegativity equalization principle gives a reasonable picture for the electronic density of molecules. The electronic densities are represented as atomic charges. Since the model gives a qualitatively correct representation of electronic densities in an atomic resolution, it is particularly useful for reactivity studies. Furthermore, it might be useful for the improvement of molecular force fields.

2. The , and polarizabilities increase rapidly with the length of the conjugated p-electron system. The components in the chain direction are largest in , and , reflecting the strong response of the delocalized p

electrons to electric fields.

3. Inclusion of heteroatoms in the p-electron system is beneficial for the , and polarizabilities, owing to either the role of additional *p* and/or *d* orbitals provided by the heteroatoms or the role of *steric* factors in the p-electron conjugation.

Acknowledgements

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