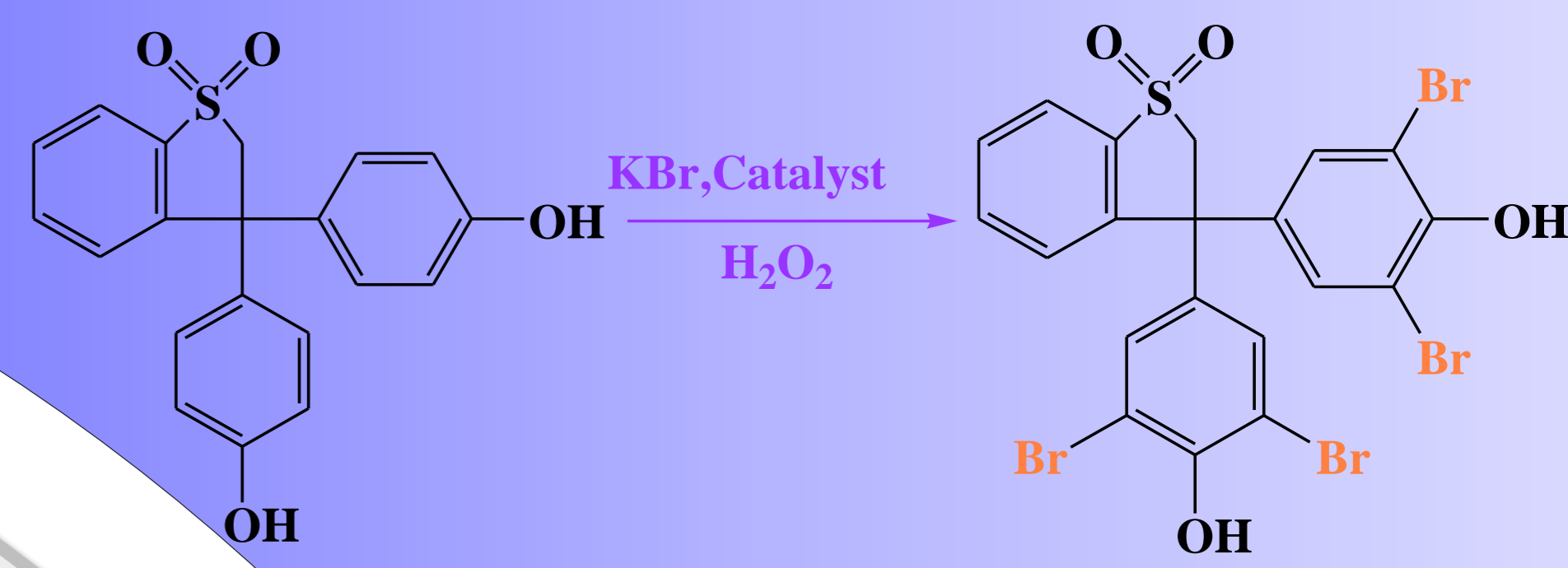


INTRODUCTION

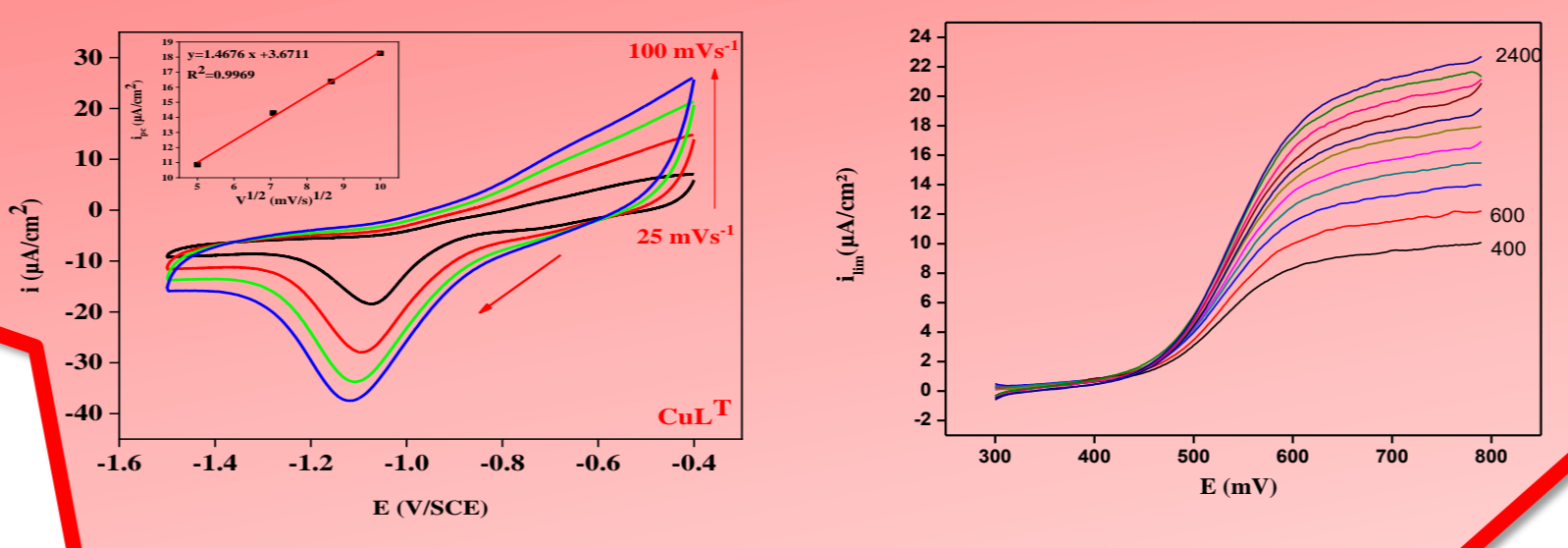
- Schiff base complexes are good candidates in catalysis oxidation reactions [1].
- Copper has received a great deal of attention due to significant roles in biochemical and catalytic properties, such as haloperoxidation, nitrogen fixation, and metalloproteins [2] [3].
- Vanadium haloperoxidases (V-HPOs) which are found in marine algae are able to accelerate the oxidative halogenation of organic compounds [4].

FUNCTIONAL MIMICS OF THE V-HPOs



Scheme. Reactive process of the bromination reaction for the complex.

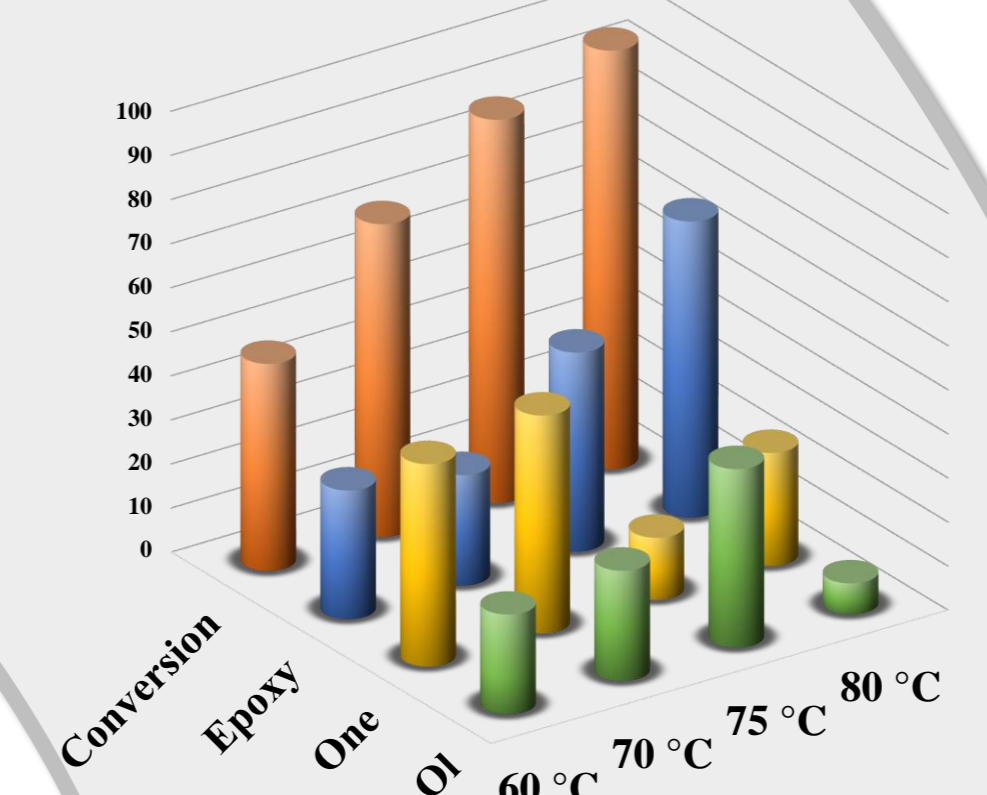
ELECTROCHEMICAL STUDY



Cyclic and Linear voltammograms of CuL^I in DMF

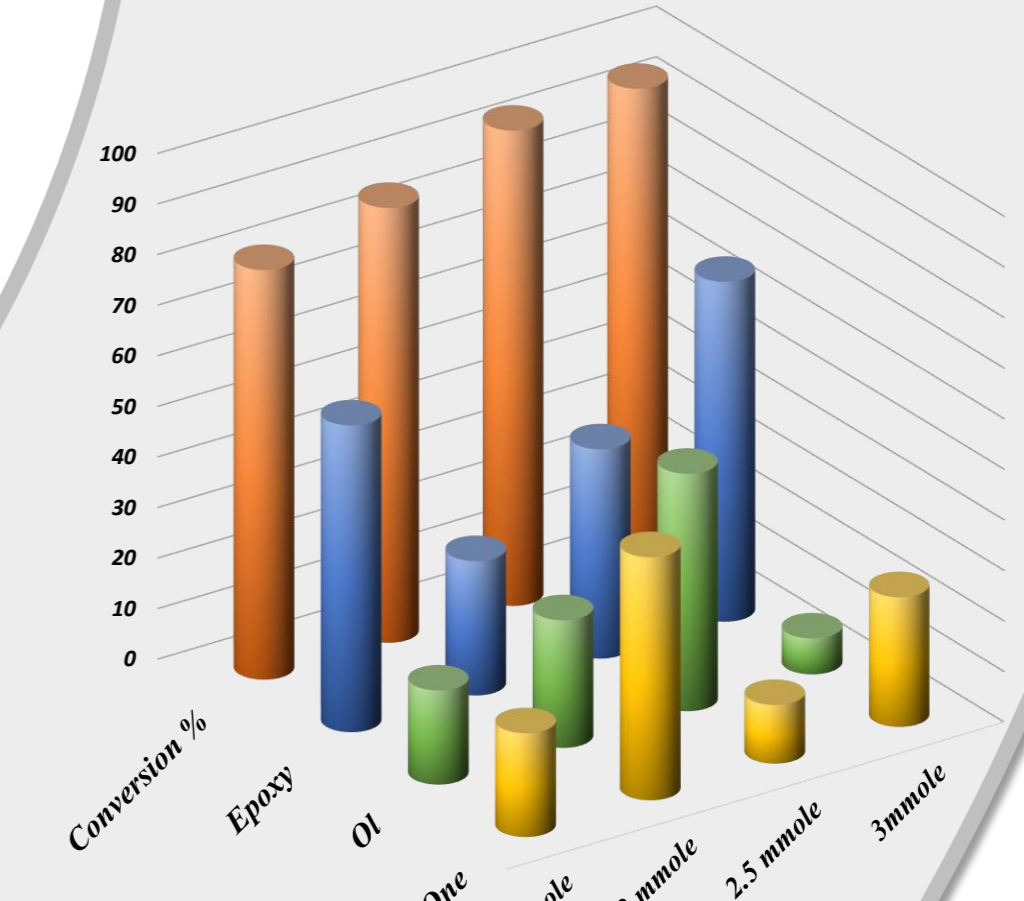
EPOXYDATION OF CYCLOHEXENE

Effect of Temperature



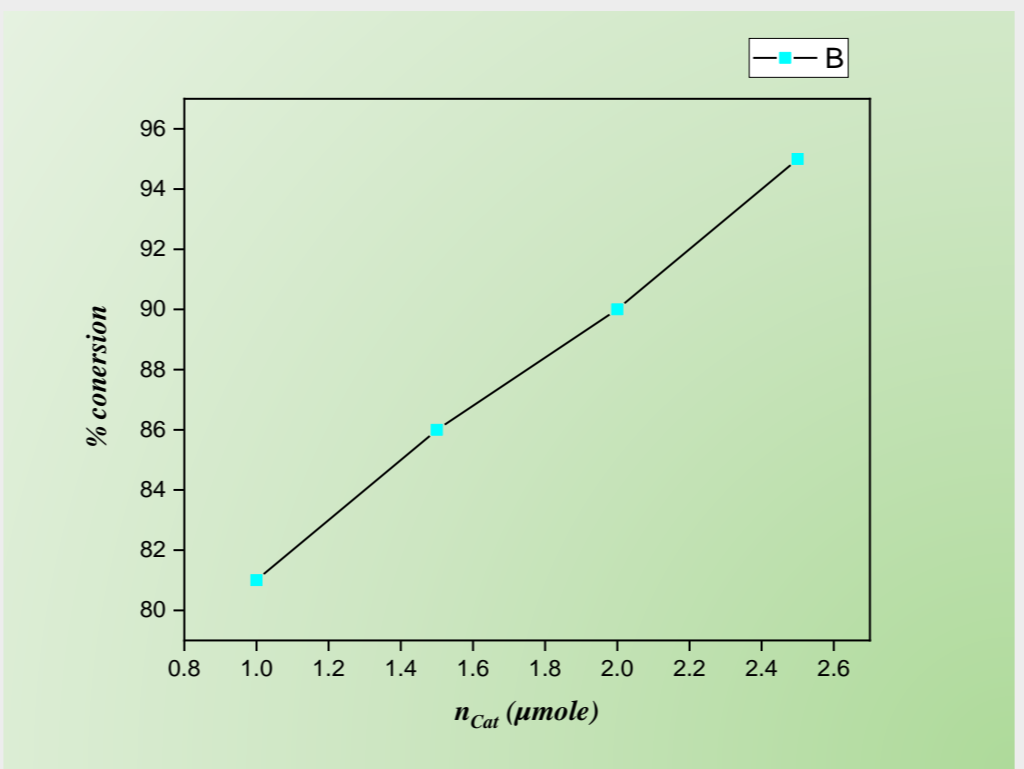
$T_{\text{opt}} = 80^\circ\text{C}$

Oxydant Effect



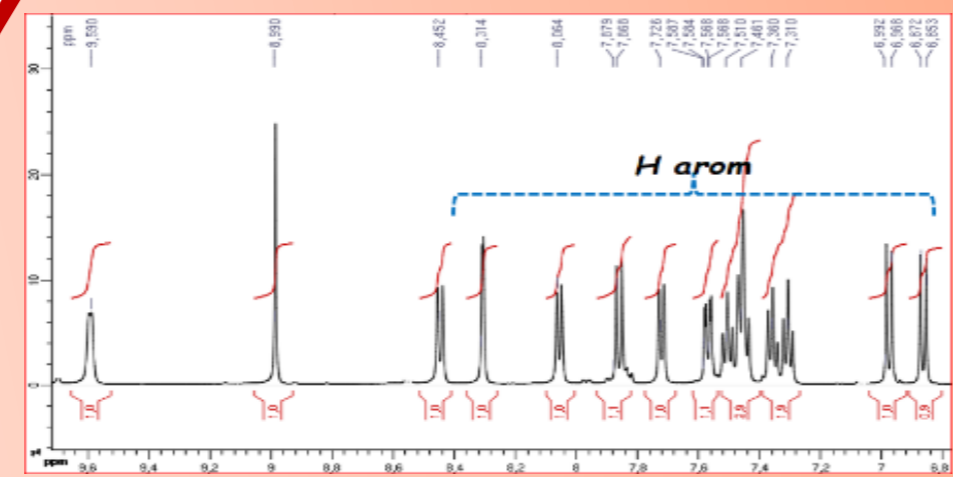
$\text{Rap}_{\text{opt}} = 1/3$

Catalyst Effect

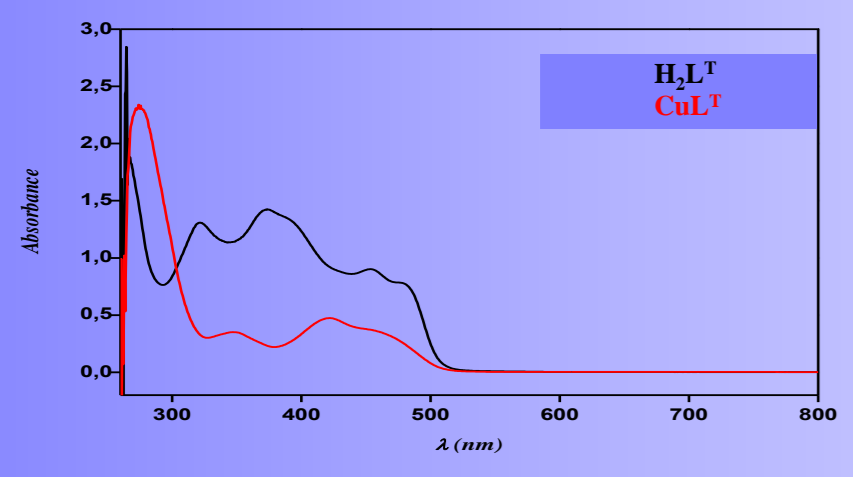


$\text{Cat}_{\text{opt}} = 1.3 \text{ mg}$

^1H NMR spectrum of H_2L^I in $\text{DMSO-d}_6/\text{TMS}$ (500MHz)



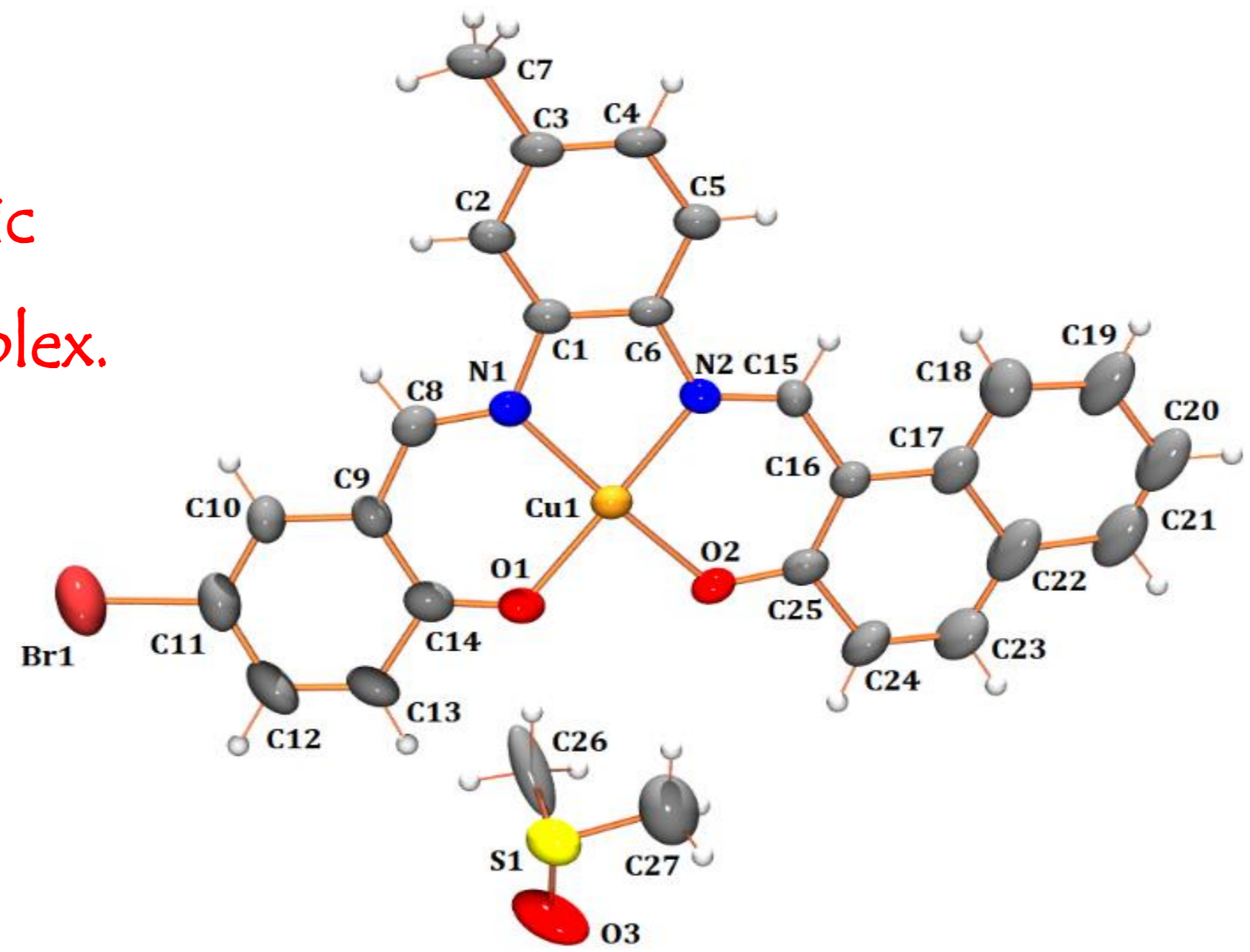
IR spectra of H_2L^I and CuL^I



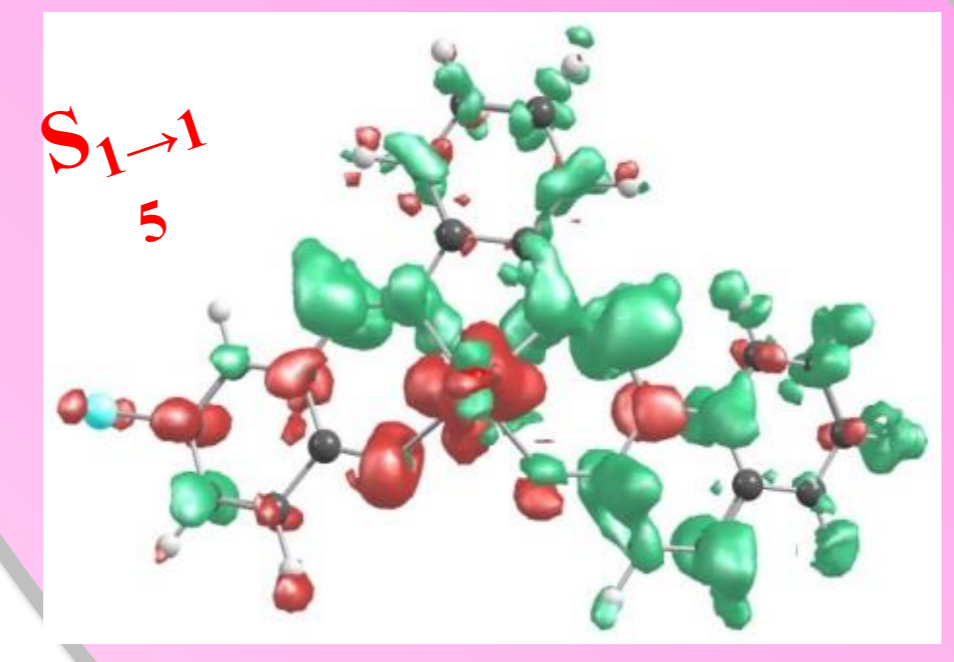
Crystal structure, Spectroscopic Studies, DFT Calculations, Cyclic Voltammetry and Catalytic Activity of a Copper Schiff Base Complex.

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DFT STUDY



$q^{CT} = 0,999$ $t = -1,373$

CONCLUSION

- The work described in this paper involves the synthesis and structural characterization of a Copper complex with a tetradentate diazomethine ligand.
- The study of the electrochemical behavior shows that the electronic transfer is controlled by diffusion.
- The objective of this work is to evaluate catalytic performances of the oxovanadium complex in bromination of phenol red and in epoxidation of cyclohexene.

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- [1] : K.C. Gupta, A.K. Sutar, C.-C. Lin, Coord. Chem. Rev. 253 (2009) 1926
- [2] : A.M.A. Alaghaz, B.A. El-Sayed, A.A. El-Henawy, R.A.A. Ammar, J. Mol. Struct. 1035 (2013) 83.
- [3] : A. Gennaro, A.A. Isse, F. Maran, J. Electroanal. Chem. 507 (2001) 124.
- [4] :C.P. Horwitz, R.W. Murray, Mol. Cryst. Liq. Cryst. 160 (1988) 389.



D. BOUCHERABINE

Figure. Oxidative bromination of phenol red catalyzed by CuL^I . Spectral changes at 10 min intervals.

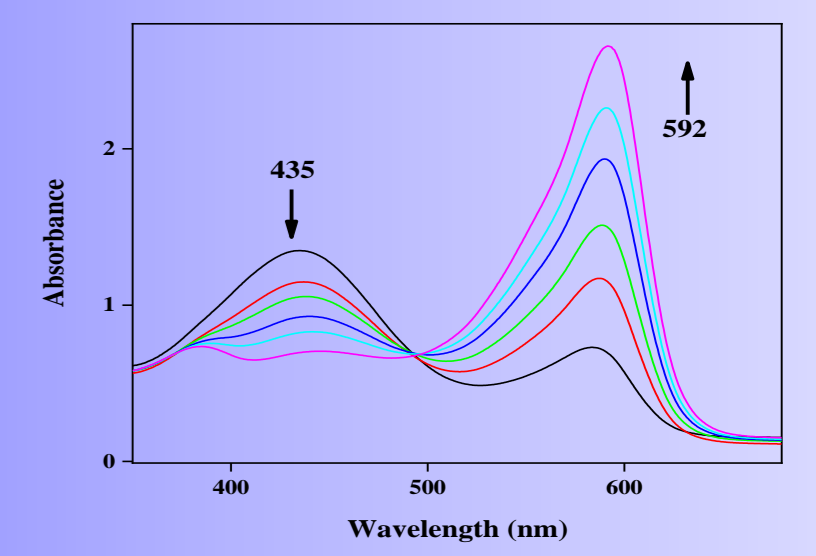
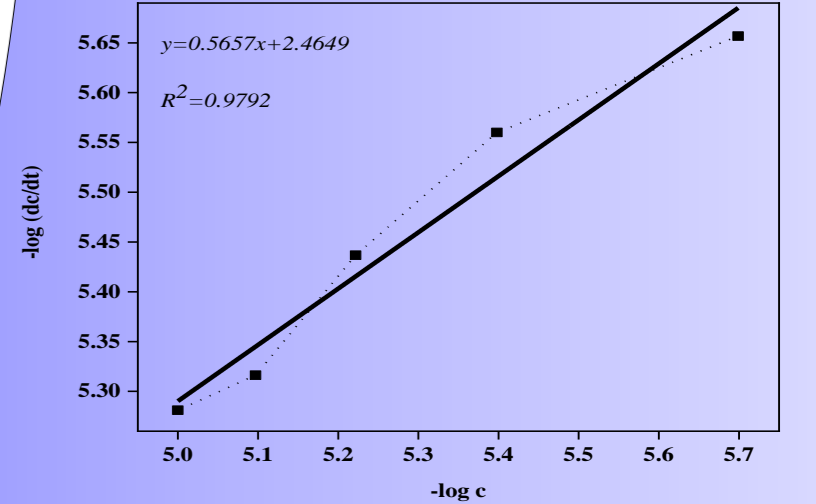


Figure. A series of linear calibration plots of the absorbance at 592 nm dependence of time for different concentration of CuL^I .



$k = 0.227 \times 10^3 (\text{M l}^{-1})^{-2} \text{s}^{-1}$