

Corrosion Behaviour of SS316 in 3.5% NaCl: A TiO₂ Thin-Film EIS Study

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INTRODUCTION & METHOD

Stainless steel SS316 is the workhorse alloy of marine, petrochemical, and biomedical industries, prized for its Cr-Ni-Mo passive oxide film. However, chloride ions aggressively penetrate and displace this passive film, triggering localized pitting and crevice corrosion that drastically reduce component service life.

Corrosion costs the global economy an estimated US\$2.5 trillion annually (3.4% of global GDP), with available best practices able to save US\$375–875 billion per year (NACE IMPACT, 2016).

Aim: To evaluate whether an RF-sputtered TiO₂ thin film can enhance the corrosion resistance of SS316 in 3.5 wt.% NaCl, using Electrochemical Impedance Spectroscopy (EIS) and linear polarization to quantify the protective effect.

Materials: SS316 (ASTM A276/A240; 17% Cr, 12% Ni, 2.5% Mo) abraded with SiC papers (120→1200 grit), cleaned, and immersed in freshly prepared 3.5 wt.% NaCl (AR grade, DI water >18 MΩ·cm).

TiO₂ Coating: Deposited by RF magnetron sputtering (~10⁻⁷ Torr base pressure, 20 mTorr working pressure, Ar/O₂ = 30/12 sccm, 12 rpm rotation) to a thickness of ~50 nm.

Electrochemistry: AUTOLAB PGSTAT302N, 3-electrode cell (SS316 WE ~1 cm², Pt mesh CE, Ag/AgCl RE, +0.197 V vs. SHE). EIS: 0.1 Hz–1 MHz, 10 mV AC. Linear polarization: ±0.1 V vs. OCP, 1 mV/s. All tests at 23±2°C, in triplicate.

TEST SETUP



Figure 1: TiO₂-coated SS316 specimen



Figure 2: AUTOLAB PGSTAT302N electrochemical cell

RESULTS & DISCUSSION

EIS Nyquist plots (Figs. 1–2) show bare SS316 impedance contracting from ~60,000 Ω (Day 2) to ~28,000 Ω (Day 4), confirming passive film breakdown, with partial re-passivation by Day 8. In contrast, TiO₂-coated SS316 exhibits $-Z'' > 130,000$ Ω on Day 1 — a 77% increase over bare SS316 at the same time point — and remains protective through Day 3.

Linear polarization (Figs. 3–4) confirms this trend quantitatively: corrosion current (*i*_{corr}) drops by 94.9% and corrosion rate falls by ~70% with the TiO₂ coating, while *E*_{corr} shifts +4 mV toward more noble behaviour — both EIS and polarization data independently agree.

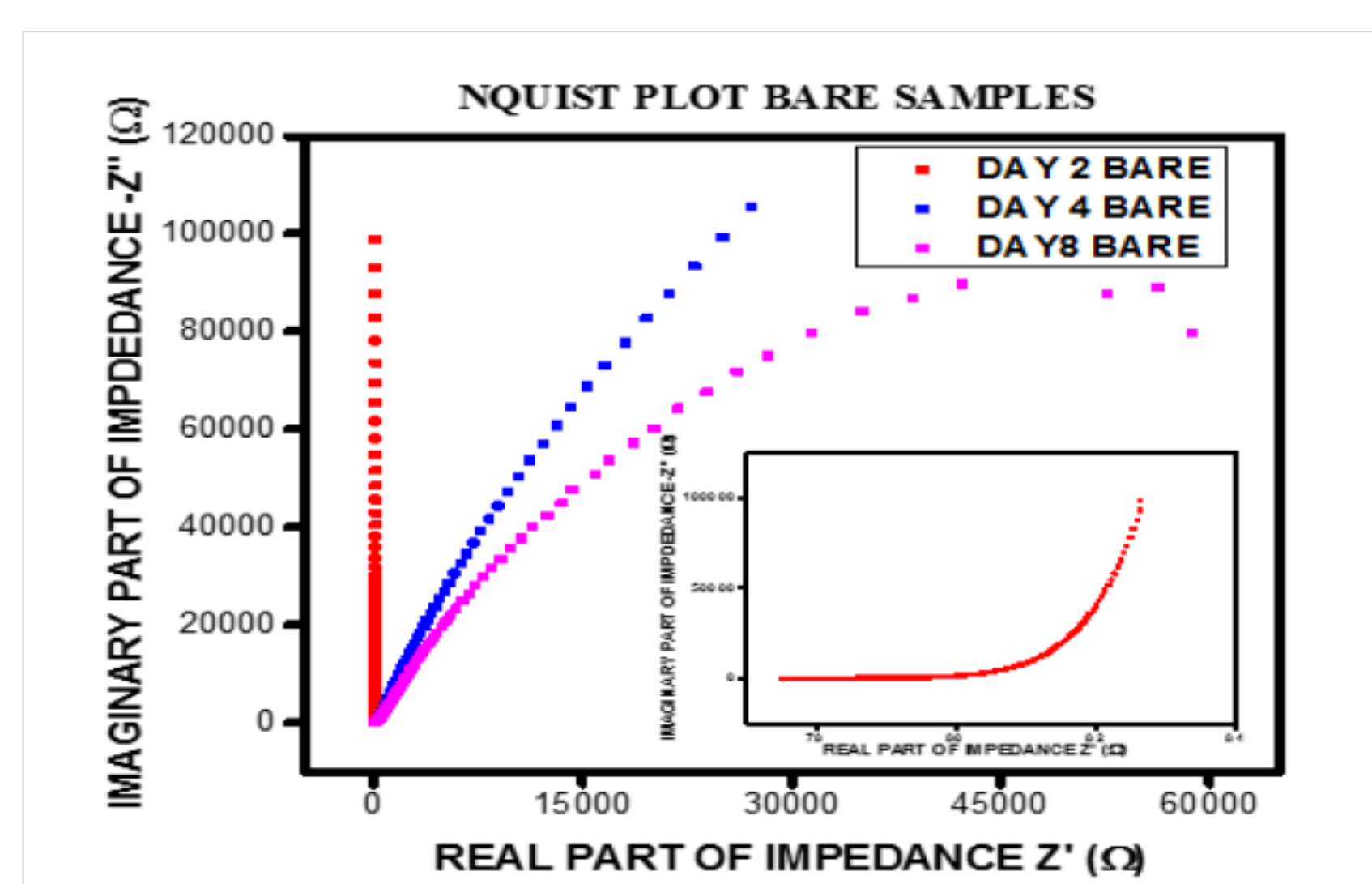


Figure 3. Nyquist plot — bare SS316 (Day 2, 4, 8). Z' contracts from ~60,000 Ω (Day 2) to ~28,000 Ω (Day 4), confirming passive film breakdown.

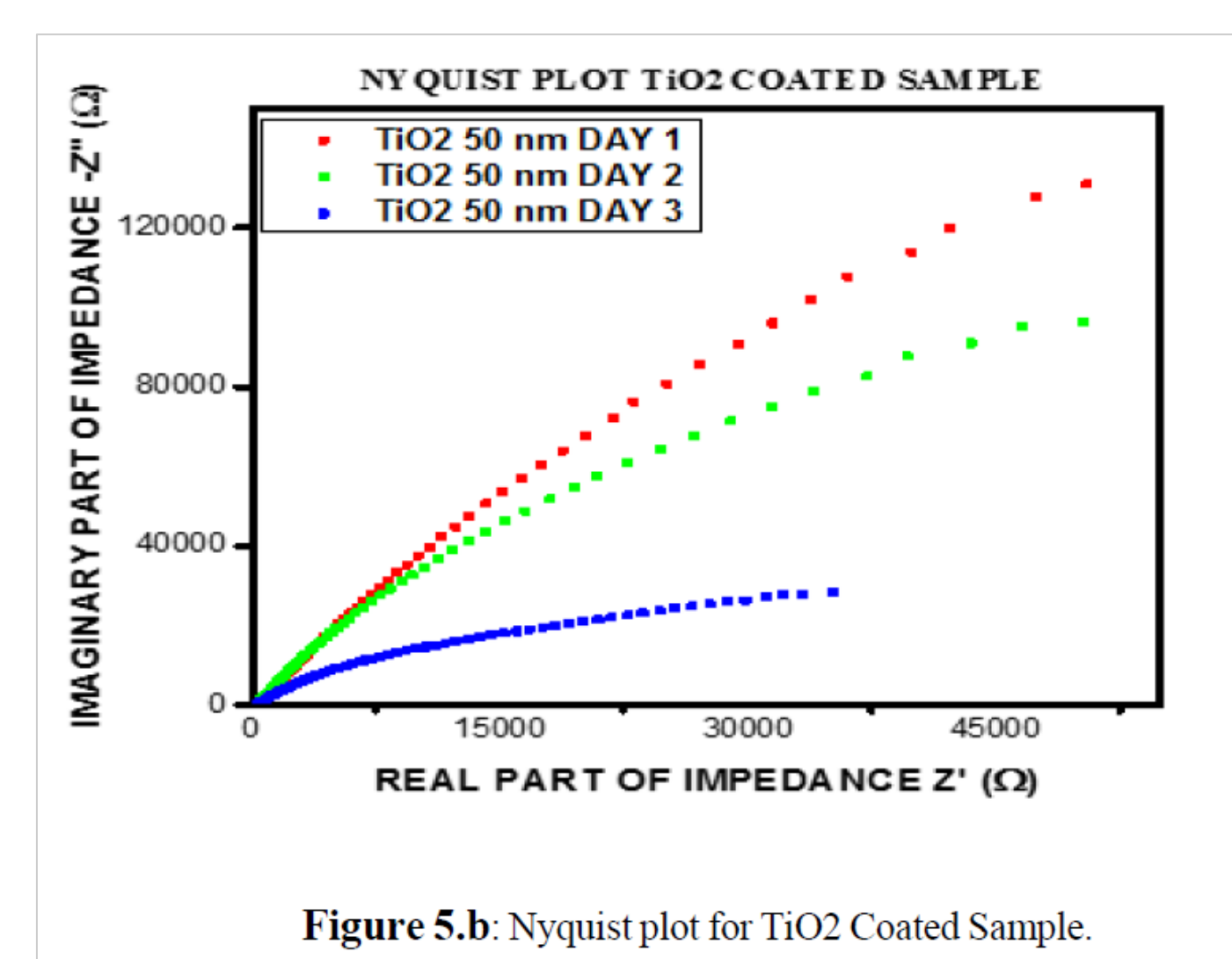


Figure 4. Nyquist plot — TiO₂-coated SS316, 50 nm (Day 1, 2, 3). $-Z''$ exceeds 130,000 Ω on Day 1 — the largest impedance recorded in this study.

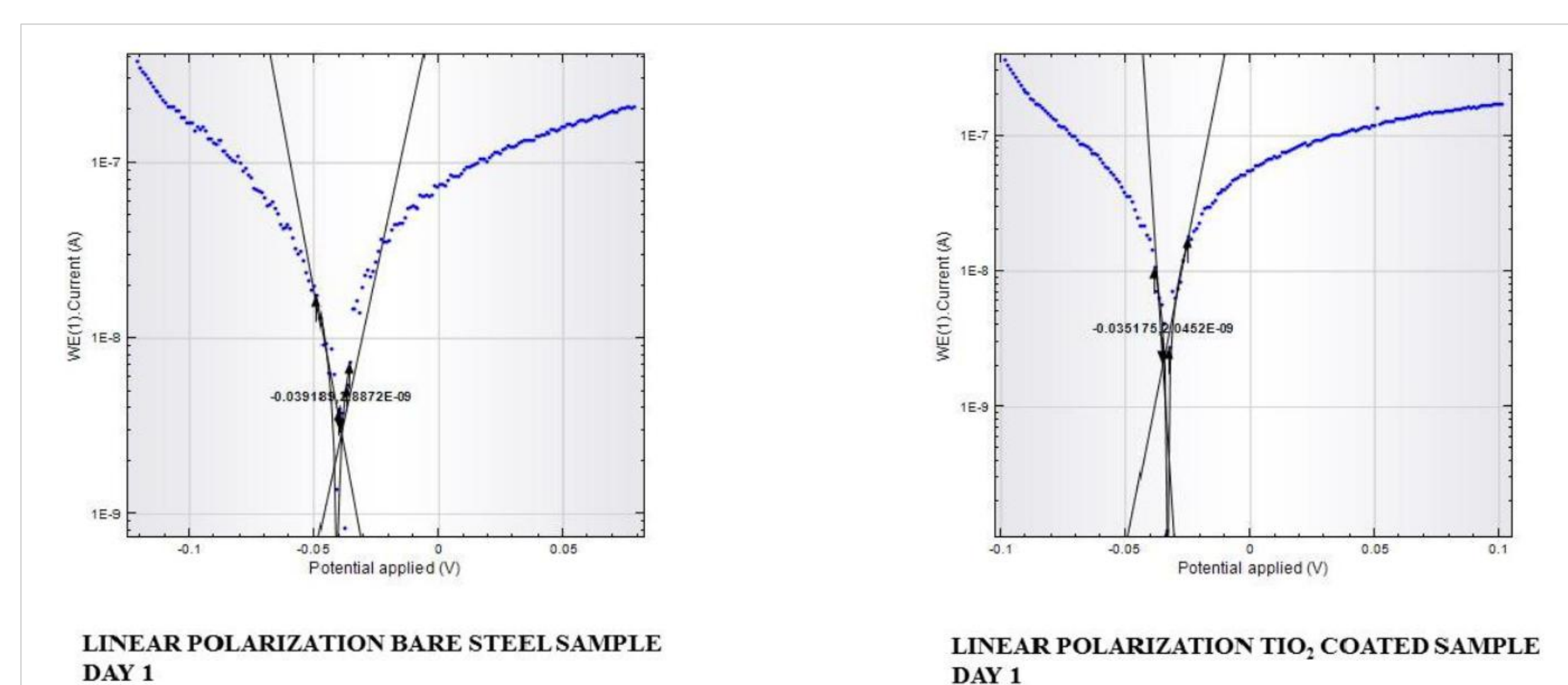


Figure 5. Linear polarization curves, Day 1. Bare: *E*_{corr} = −0.039 V, *i*_{corr} = 0.889 nA. TiO₂-coated: *E*_{corr} = −0.035 V, *i*_{corr} = 0.0452 nA (−94.9%).

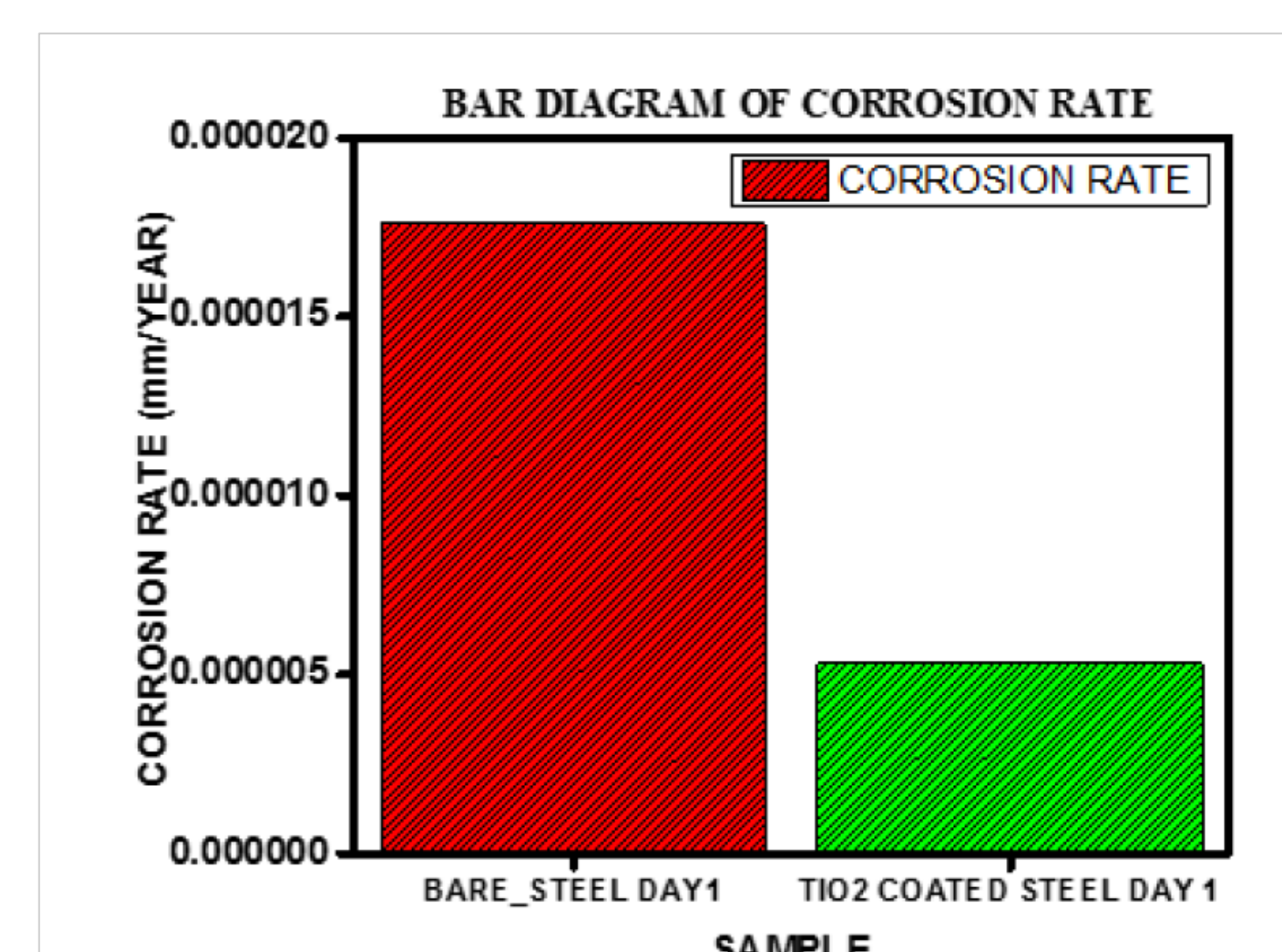


Figure 6. Corrosion rate comparison, Day 1. TiO₂ coating reduces corrosion rate by ~70% ($1.77 \times 10^{-5} \rightarrow 5.3 \times 10^{-6}$ mm/year).

E_{corr} Shift
+4 mV

−«0.039» V → −0.035 V

***i*_{corr} Reduction**
−94.9%

0.889 nA → 0.0452 nA

Corrosion Rate
−70.1%

$1.77 \times 10^{-5} \rightarrow 5.3 \times 10^{-6}$ mm/yr

EIS Impedance (Day 1)
+77%

75,000 Ω → 133,000 Ω

CONCLUSIONS

- RF-sputtered TiO₂ thin films (~50 nm) significantly enhance the corrosion resistance of SS316 in 3.5 wt.% NaCl, confirmed independently by both EIS and linear polarization.
- Corrosion rate reduced by ~70% ($1.77 \times 10^{-5} \rightarrow 5.3 \times 10^{-6}$ mm/year); *i*_{corr} reduced by 94.9%; EIS impedance increased by 77% at Day 1.
- The protective effect is attributed to a stable, compact TiO₂ barrier layer that limits electrolyte access and suppresses charge-transfer reactions — directly relevant to marine, petrochemical, biomedical, and desalination industries.

FUTURE WORK/ REFERENCES/ACKNOWLEDGMENT

Future Work: XRD/SEM/XPS characterisation; 50 nm vs. 100 nm thickness comparison; CPE equivalent circuit modelling; extended immersion studies.

Selected References: 1. Koch et al., NACE IMPACT Study, 2016. 2. Pan, Thierry & Leygraf, J. Biomed. Mater. Res., 1997. 3. González-Torres et al., Surf. Coat. Technol., 2023. 4. Santos et al., Steel Res. Int., 2024. 5. Revilla et al., Electrochim. Acta, 2020.

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