

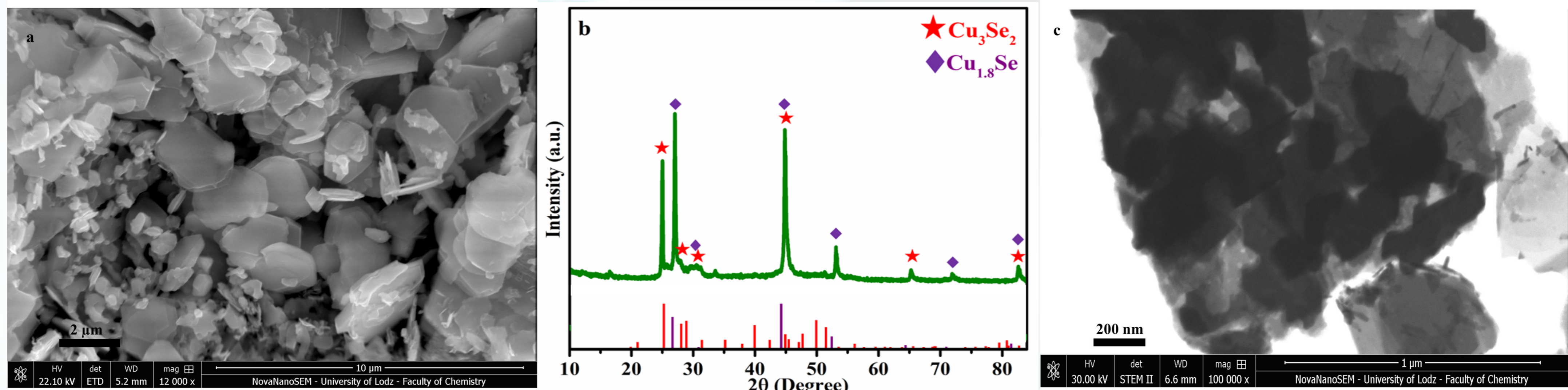
# Synthesis of Copper Selenide Nanobelts for Electrochemical Detection of Hydrogen Peroxide

Nabi Ullah, Saraf Khan, Dariusz Guziejewski,

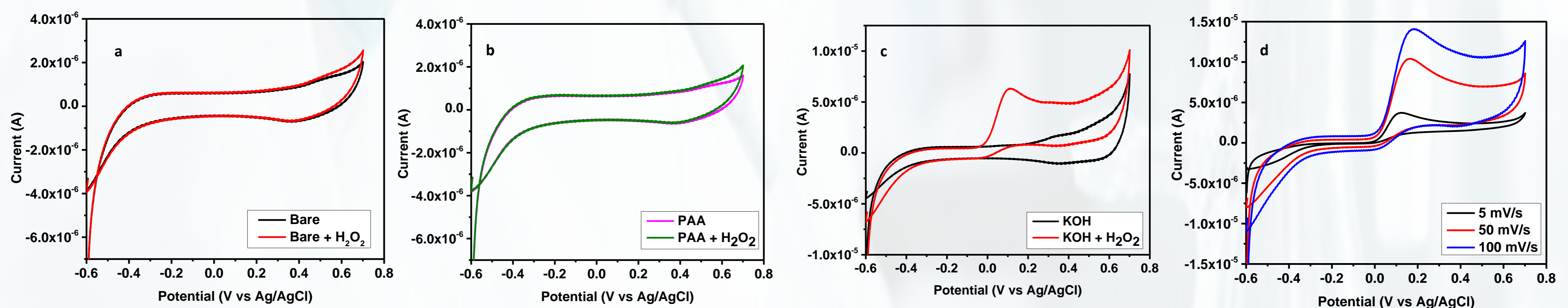
Departament of Inorganic and Analytical Chemistry Faculty of Chemistry, University of Lodz  
Pomorska 163, 92-235, Poland

## Abstract

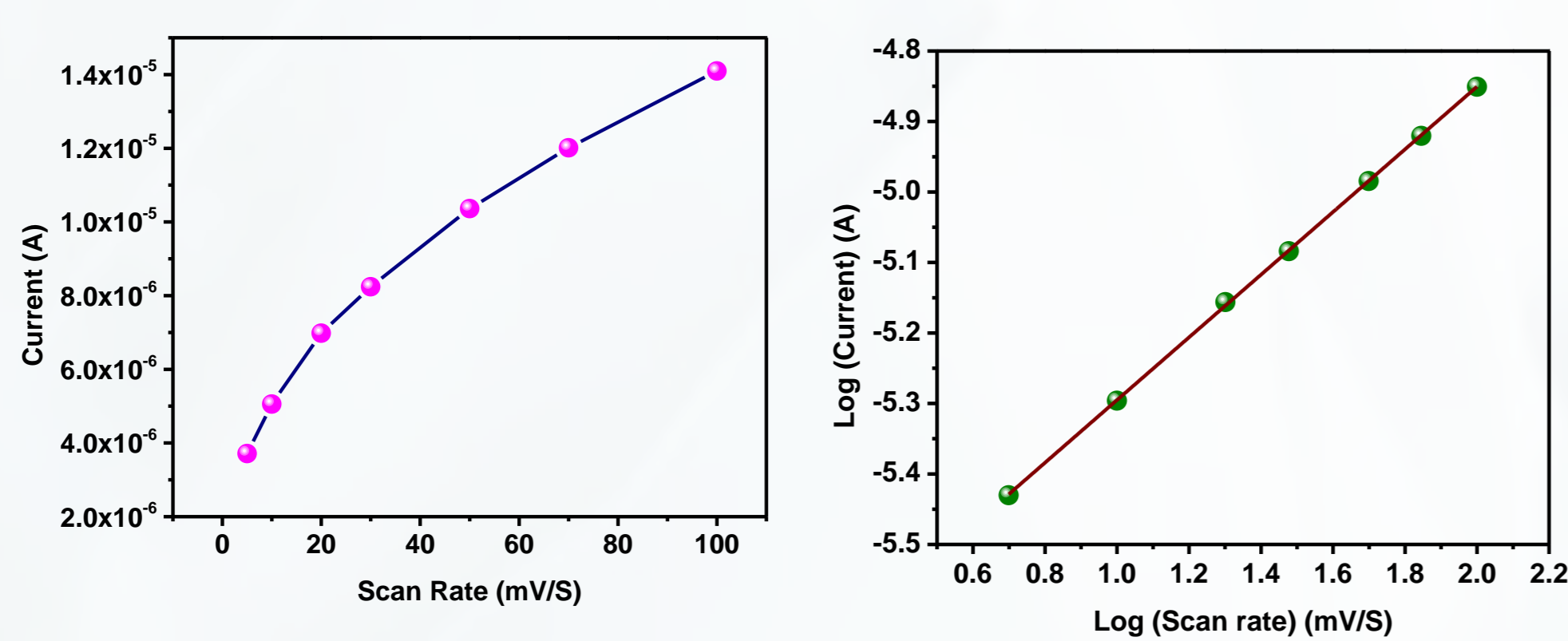
Hydrogen peroxide is a byproduct of many physiological reactions in living organism. Thus, its concentration can help determine certain diseases, for example lung cancer or asthma. [1, 2]. Additionally, its detection is in the center of interest for security reason, as it is part of homemade peroxy-explosive materials [3, 4]. copper selenide was synthesized using simple solvothermal method with nanobelt morphology, which was characterized by XRD, SEM and TEM. Copper selenide, as an electrocatalyst, was used for hydrogen peroxide determination in electrolyte of 0.1M KOH, and 0.25% PAA, covering a detection range from 0.03 to 1 mmol. The catalyst showed LOQ ( $2.49 \times 10^{-5}$  mol/L) and LOD ( $7.48 \times 10^{-6}$  mol/L) for hydrogen peroxide detection. Under optimized reaction conditions in SWV, it exhibited LOQ and LOD of  $2.75 \times 10^{-5}$  and  $8.18 \times 10^{-6}$  mol/L, respectively.



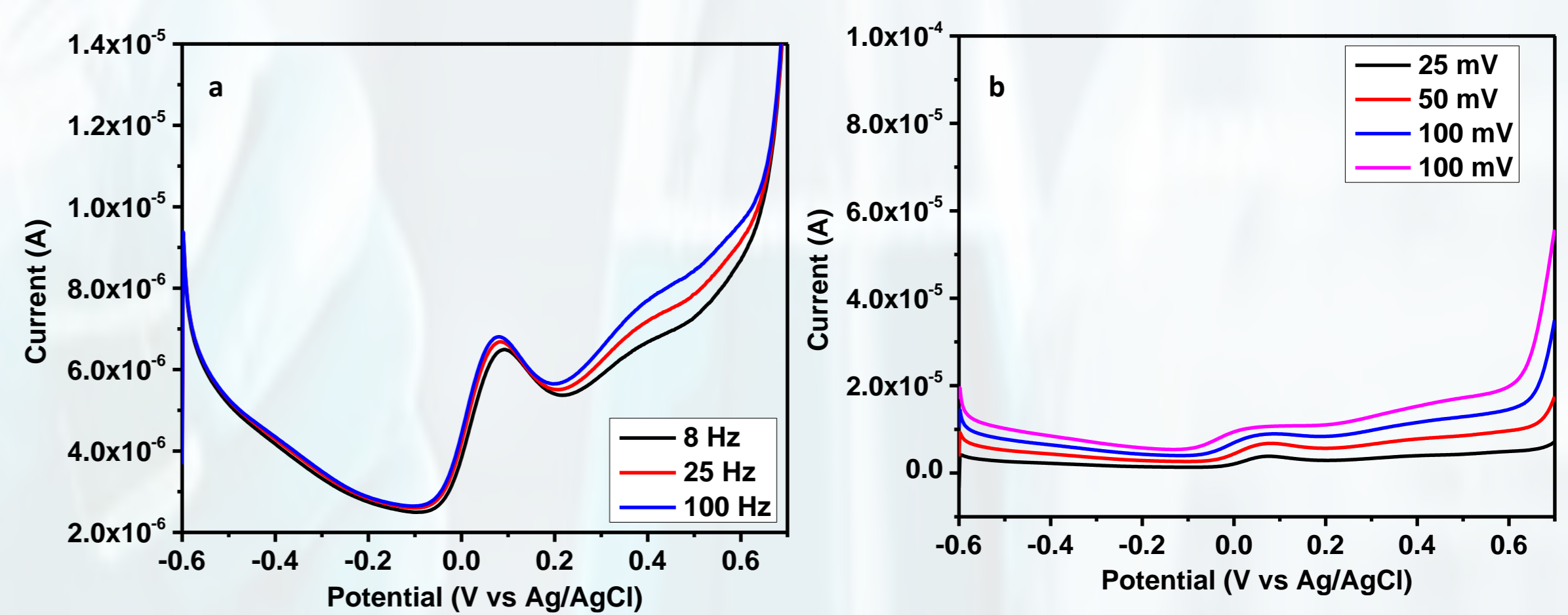
(a) SEM, (b) XRD, and (c) TEM data for copper selenide nanobelts.



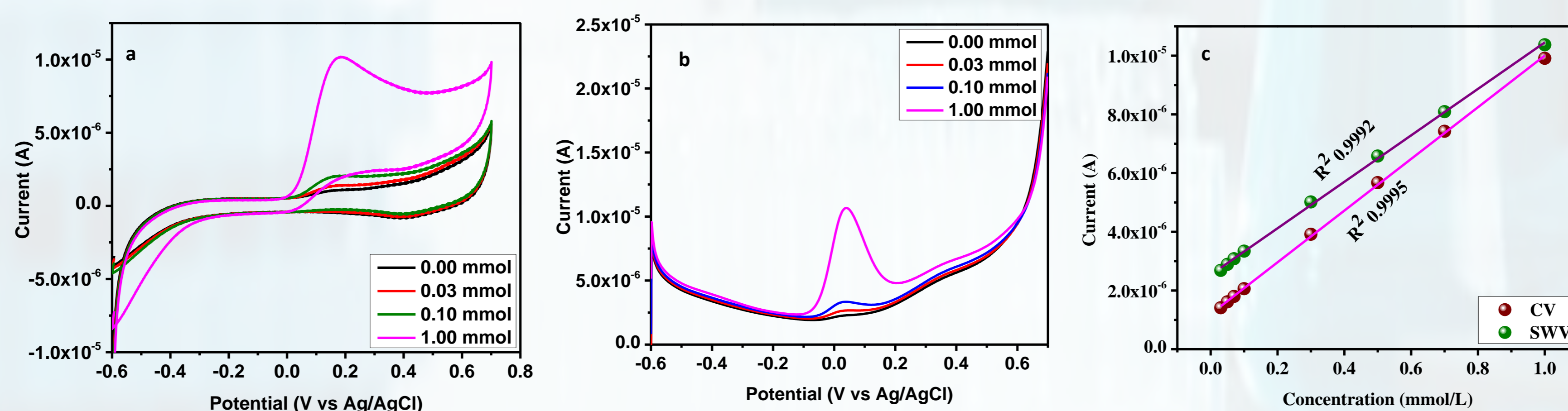
Effect of (a) bare electrode, (b) PAA, (c) KOH (supporting electrolyte), and Scan rate on hydrogen peroxide detection (oxidation).



Scan rate and current peak relation confirmed diffusion control mechanism (Slope = 0.44)



Effect of (a) Frequency (b) amplitude on H<sub>2</sub>O<sub>2</sub> detection in SWV



Detection of hydrogen peroxide under optimize condition via (a) CV, (b) SWV, and its calibration line

## Conclusion

copper selenide was synthesized using simple solvothermal method with nanobelt morphology. Copper selenide, as an electrocatalyst, was used for hydrogen peroxide determination in electrolyte of 0.1M KOH, and 0.25% PAA, covering a detection range from 0.03 to 1 mmol. The catalyst showed LOQ ( $2.49 \times 10^{-5}$  mol/L) and LOD ( $7.48 \times 10^{-6}$  mol/L) with  $R^2$  value of 0.9995 for hydrogen peroxide detection. Under optimized reaction conditions in SWV, it exhibited LOQ and LOD of  $2.75 \times 10^{-5}$  and  $8.18 \times 10^{-6}$  mol/L with  $R^2$  value of 0.9992, respectively.

## References

- [1] Z.M. Sheng, H. Huang, R.L. Niu, Z.W. Han, R.P. Jia, Sens. Actuators B: Chem., 2020, 305, 127550.
- [2] J. He, W. Zhou, J. Sunarso, X. Xu, Y. Zhong, Z. Shao, X. Chen, H. Zhu, Electrochim. Acta, 2018, 260, 372-383.
- [3] Q. Chen, L. Yang, K. Guo, J. Yang, J.-M. Han, ACS Omega, 2021, 6, 11447-1145
- [4] L. Dunn, H.S.A. Al Obaidly, S.E. Khalil, Forensic Chem., 2018, 10, 5-14.