

# Electro Synthesis and Characterization of PANI and PANI/ZnO Composites Films

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**Abstract:** The polyaniline (PANI) and ZnO doped polyaniline composite (PANI/ZnO) films were synthesized by electrochemical polymerization, on transparent ITO glass substrate using cyclic voltammetry for 10th cycle by potential sweep from  $-200$  to  $900$  mV vs. SCE in sulfuric acid medium with aniline monomer precursor and  $\text{LiClO}_4$  supporting salt at room temperature. The doping effects of ZnO on PANI properties were studied; electrical properties of PANI and PANI/ZnO composite were estimated from Mott-Schottky analysis which demonstrated an n-type conductivity for all films with carrier concentrations in order of  $10^{17}$  and  $10^{19}$  for PANI and PANI/ZnO composite respectively. The effect of the concentration of ZnO dopant in the electrolyte on the morphology and the composition of the PANI were investigated from scanning electron microscopy coupled Energy Dispersive Spectroscopy analysis (SEM+EDS). UV-VIS was investigated to study the optical characteristics of the films such as transmittance and band gap energy. The characteristics of the films were found to be depended to the concentration of ZnO particles in the electrolyte during electrochemical polymerization. Fourier transform Infrared (FTIR) confirms the links and the interaction between PANI and ZnO dopant.

**Keywords:** polyaniline; ZnO dopant; voltammetry cyclique; semiconductor

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## 1. Introduction

The  $\pi$ -conjugated polymers have currently received much attention of scientists and researchers include polypyrrole, polyfurane and especially polyaniline which were investigated in the literature [1–3]. Because of their various potential applications in the fields of batteries, electro chromic display devices, sensors, and organic electrical devices [4–10]. The polyaniline is one of conjugated polymers which has been widely studied due to its easy synthesis and reversible acid–base chemistry in aqueous solution, good chemical and thermal stabilities, redox reversibility, its metal behavior and high electric conductivity when it was doped in acid medium [11,13].

In recent years, the doping of conjugated polymers with different inorganic semiconducting nanoparticles such as  $\text{SnO}_2$ ,  $\text{TiO}_2$  and  $\text{ZnO}$ ...etc to formed composite materials has widely studied for optoelectronic applications [14,16]. This doping has prove an amelioration in many properties of the composites like great enhancement in PL property of PANI/  $\text{ZnO}$  composite [17] and increasing in conductivity of polyaniline/zinc oxide or polyaniline/cadmium oxide nanocomposite compared with polyaniline [22,24]. Various methods and ways of syntheses were investigated to fabricate the polyaniline and the polyaniline/semiconductor composite, for example Mostefaie and al. [21] have prepared these materials by an in situ chemical oxidative method of aniline monomers in the presence of  $\text{ZnO}$  nanorods, a novel Pickering emulsion route was reported by He to prepear a

polyaniline/nano-ZnO composite fibers [11], the solution casting technique was utilized for preparation of polyaniline (PANI)/ZnO nanocomposites film in the study of Ahmed and al. and that of Singla and al for synthesis of polymer/ZnO nanoparticles composite [22,23], Patil and al. have prepared the polyaniline/ZnO (50%) nanocomposites from spin coating method on glass substrates [24,25], also the electrochemical polymerization method is used more often because it is environmentally friendly and presents several advantages [26]. Many current works have been taken this later as synthesis method of the polymers and the composite polymer/semiconductor like PANI, PANI/TiO<sub>2</sub>, polyaniline/LiMn<sub>2</sub>O<sub>4</sub> [3,27].

The electrical properties of the conducting polymers and composites were the key parameters to evaluate the importance and the utility of these materials. Different techniques were utilized to investigate these characteristics (conductivity, resistivity and carriers concentration) like Hall effect measurement, four point probe, but the Mott-Schottky analysis has been not reported yet, despite its simplicity and importance of the parameters derived from this technique like carriers concentration and flat band potential which were characteristics of organic and inorganic semiconductors.

This work involves the electrochemical polymerization of polyaniline (PANI) and the polyaniline/zinc oxide composite (PANI/ZnO) films from the Aniline (ANI) monomer and Zinc oxide particles in sulfuric acid medium with LiClO<sub>4</sub> as supporting electrolyte on ITO glass electrodes. Cyclic voltammetry was investigated as electrochemical technique for synthesis, electrical properties of the PANI and the PANI/ZnO were extracted from Mott-Schottky analysis. In other hand the optical characteristics were studied using UV-VIS and FTIR spectroscopy.

## 2. Experiments Procedures

### 2.1. materials and Instrumentation

The electrodeposition of PANI is performed in aqueous acid solutions of H<sub>2</sub>SO<sub>4</sub>, the supporting electrolyte used was lithium perchlorate (LiClO<sub>4</sub>) which is a pure salt for analysis and soluble in organic and aqueous solutions, and because of its electrochemical stability on a large domain of potential. The ZnO 99%, a powder used as a doping semiconductor, and the monomer was aniline (An) with 99.99% purity. Distilled water was for the preparation of electrolyte solution in all stages. All the electrochemical measurements (cyclic voltammetry and Mott-Schottky analysis) were performed using a Voltalab 402 type PGZ from Radio meter Analytical, coupled with a computer equipped with a software (voltmaster 4) to select the electrochemical technique and fix the desired parameters. A three-electrodes cell was used to carry out the electrochemical experiments; a platinum wire (diameter 2 mm) was used as a working electrode for electrochemical characterization. For UV-VIS and Scanning electron microscopy (SEM) coupled, ITO coated glass electrode were used as working electrodes (area of 1 cm<sup>2</sup>), a platinum wire served as an auxiliary electrode and a saturated calomel electrode (SCE) as reference, all the potential values were referred to this electrode.

The optical transmittance measurements were performed on a Shimadzu UV-spectrophotometer UV-1800 coupled with UV Probe software which works in the range of 200 to 1100 nm. The FTIR analysis of PANI and PANI/ ZnO composite were performed by Jasco-FT/IR-4200 spectrophotometer. SEM and EDX measurements were carried out with a Quanta 250 with tungsten filament from FEI Company.

### 2.2. Preparation of the PANI and (PANI/ZnO) Composite Film

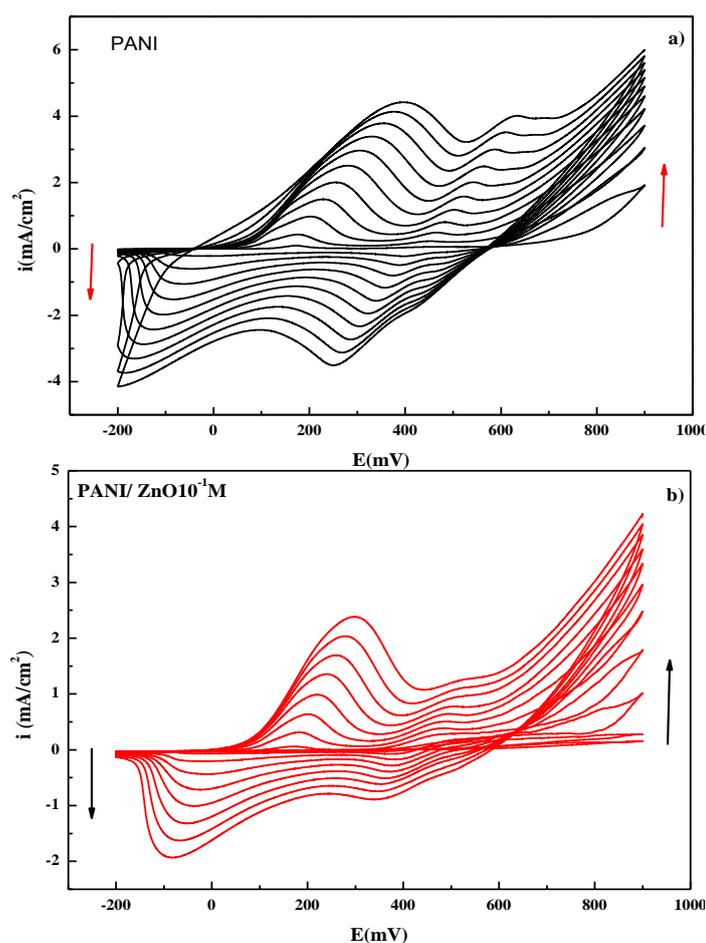
The smoothing pretreatment of the surface Pt electrode was preceded by mechanical polishing, using emery paper down to 1200 grade and degreasing with acetone/ethanol mixture, washing with distilled water and drying. The ITO substrates were ultrasonically cleaned in methanol, acetone and distilled water for 10 min and etched in nitric acid 45% for 2 min to activate the surface of the electrode. The electro polymerization of aniline in the absence and the presence of ZnO with different concentrations (10<sup>-4</sup>, 10<sup>-3</sup>, 10<sup>-1</sup> M) was performed in potentiodynamic conditions in a cell containing 1M H<sub>2</sub>SO<sub>4</sub> + 0.1M LiClO<sub>4</sub> and 10<sup>-1</sup> M aniline. The polyaniline (PANI) and the composite material

(PANI/ZnO) were obtained on the electrode by cyclic voltammetry (cycling) in the potential range between  $-0.2$  and  $0.9$  V/SCE, at a scan rate of  $10$  mV/s. Before each experiment, the electrolytic solution was degassed by argon bubbling for  $15$  min.

### 3. Results and Discussion

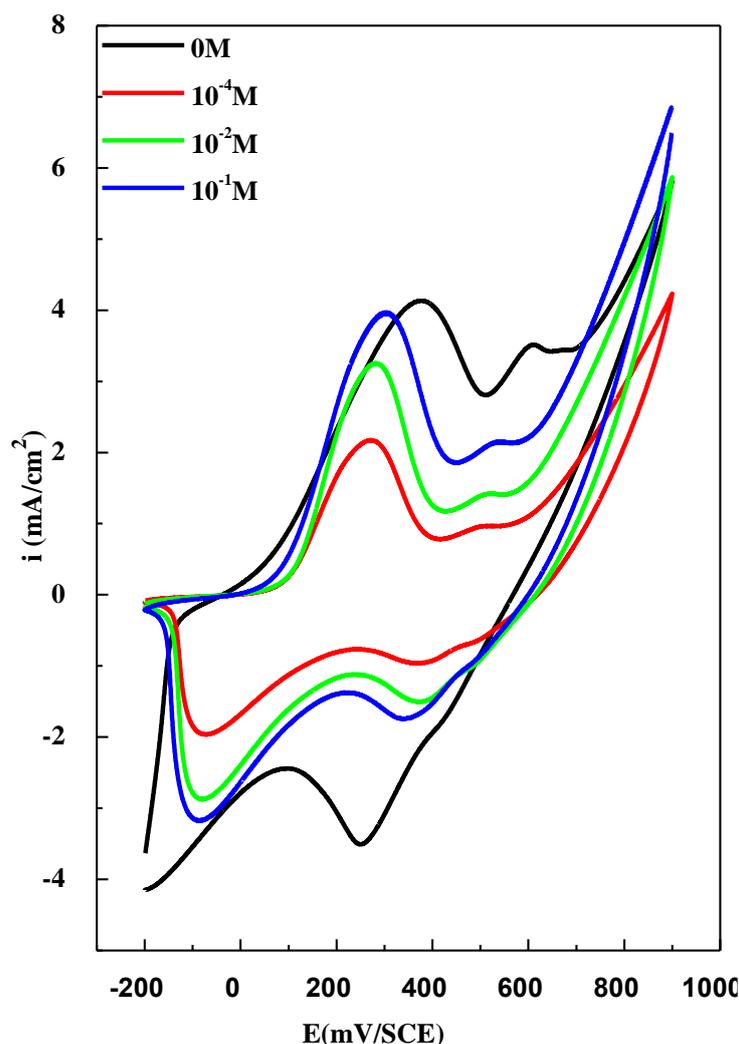
#### 3.1. Cyclic Voltammetry

Figure 1 shows a cyclic voltammograms (CV) corresponding to the electrochemical characterization of PANI (a) and the composite material based on zinc oxide particles PANI/ZnO (b) on ITO glass substrate with  $1$  cm<sup>2</sup>. These voltammograms were recorded in the potential range between  $-0.2$  and  $0.9$  V/SCE, and with scan rate of  $10$  mV s<sup>-1</sup>. On the first cycle, The cyclic voltammograms show during the positive potential scan three anodic peaks which are, respectively, observed in the potential  $E_{pa1} = 0.174$ ,  $E_{pa2} = 0.468$ , and  $E_{pa3} = 0.897$  V/SCE and during the negative scan potential three cathodic peaks that appear two at  $E_{pc1} = -0.024$ ,  $E_{pc2} = 0.387$ , and  $E_{pc3} = 0.553$  V/SCE, which correspond to different oxidation and reduction states of the PANI. The first redox peak corresponds to the transition of the leucoemeraldine (LE) to emeraldine (EM) on the forward scan, and on the backward scan switching back to LE can be seen. The third redox peak corresponds to the transition of EM to pernigraniline (PG), while on the reversed scan the PE reduction to EM is observed. It was established that the second pair may have a different origin: it can indicate the presence of oligomers or degradation products, [28,29].



**Figure 1.** Electrochemical polymerization of: (a) ANI with 10 cycles, (b) ANI in presence of  $10^{-1}$  M ZnO with 10 cycles, in  $1$  M  $H_2SO_4 + 10^{-1}$  M  $LiClO_4 + 10^{-1}$  M ANI with scan rate of  $10$  mV/s.

After the second cycle, the current intensity of the oxidation and reduction peaks progressively increased with the number of cycles, indicating the formation and the growth of the conducting polymer film and suggesting a systematic increase in the electrode area as a result of the actual deposition of PANI [30,31]. Comparing to the CV of the composite material PANI/ZnO with  $10^{-1}$  M of ZnO particle in the electrolyte, we observed a low current intensity of the redox peak, this can be due to the dispersion of ZnO nanoparticles in polymer matrix, and as a consequence the decrease of active surface. The main parameter in our experiments was the effect of the dispersed ZnO concentration in the electrolyte, which varied in this extent of concentration ( $10^{-4}$ ,  $10^{-3}$ ,  $10^{-1}$  M). These quantities are added to the solution containing the Aniline monomer. Then, we plotted the first voltammogram for each concentration and the corresponding voltammograms are grouped in Figure 2.



**Figure 2.** Cyclic voltammograms relating to the deposition of PANI and PANI/ ZnO composite film on ITO substrate in electrolyte containing 1M  $\text{H}_2\text{SO}_4$ + 0.1M Aniline + 0.1M  $\text{LiClO}_4$  with different concentration of ZnO particle in electrolyte at scan rate of 10 mV/s.

As shown in Figure 2, the formation of composite PANI/ZnO, a same style of voltammograms were obtained for different concentration ZnO, but with a decrease in the current intensity of cathodic and the anodic peaks. A slight decrease was observed in the presence of  $10^{-1}$  M of ZnO, but it was noticeable for  $10^{-2}$  and especially for  $10^{-4}$  M concentration of ZnO particles. The phenomena was accompanied by a slight shift of the cathodic peak potential towards more positive values, which can be attributed to the modification of the surface electrode by the incorporation of the ZnO particles in PANI. So dispersed ZnO influences the electrochemical deposition and properties of polyaniline.

### 3.2. Mott-Schottky Analysis

The Mott-Schottky analysis was benefits to access at the electrical characteristics of semi-conductors like conductivity type, carrier's concentrations and flat band potential. This technique was based to study the comportment of semiconductor/ electrolyte junction under application of bias potential and the measuring the variation of the capacitance as a function of applied potential across the space charge layer according to the Mott Schottky equation for n type semiconductor [32,33].

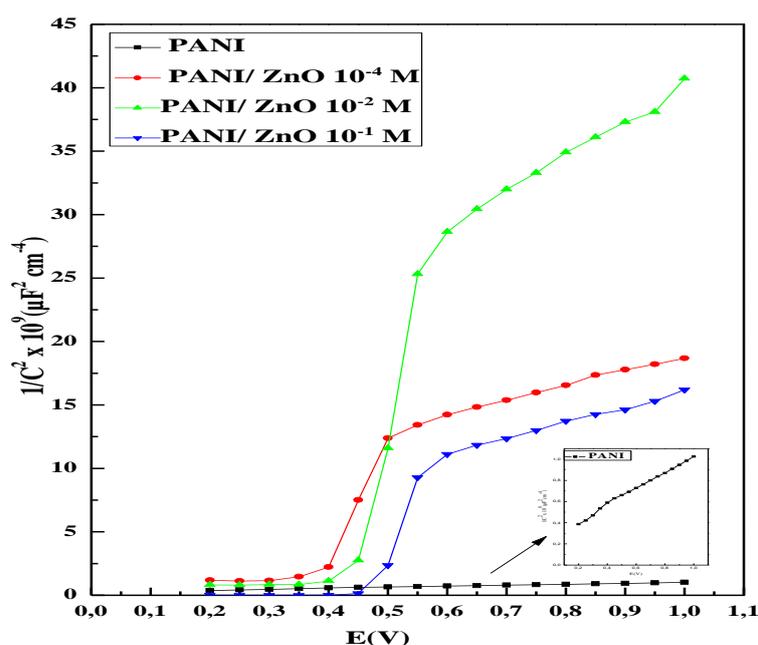
$$\frac{1}{C_{SC}^2} = 2/(e\epsilon\epsilon_0 A^2 N_D) [V - V_{fb} - \frac{KT}{e}] \quad (1)$$

where  $C_{sc}$  is the capacity of space charge region,  $\epsilon_0$  is the vacuum permittivity ( $\epsilon_0 = 8.85 \times 10^{-14} \text{ F cm}^{-1}$ ),  $\epsilon$  is the dielectric constant of the PANI layer ( $\epsilon = 674.34$ ) [34],  $N_D$  is the carrier concentration density,  $V$  is the applied potential,  $V_{fb}$  is flat band potential,  $T$  is the absolute temperature ( $T = 294 \text{ K}$ ) and  $K$  is the Boltzman constant.

Figure 3 illustrates the Mott Schottky plots of PANI films and PANI/ZnO composite films electropolymerized at various concentrations of ZnO particles contact in  $1\text{M H}_2\text{SO}_4 + 0.1\text{M LiClO}_4$  as supporting electrolyte with 500 Hz frequency, The positive slope in these plots indicates that all the samples were a n type semiconductor [35], the estimated electron densities ( $N_D$ ) were calculated from the slope of the Mott Schottky plots ( $2/e\epsilon\epsilon_0 A^2 N_D$ ). The Table 1 regroupes the electron densities values as a function of concentration of ZnO particles, we can observed that the presence of ZnO particle in PANI matrix ameliorate the carrier concentration values. This later was found to depend on the ZnO concentration and was maximal for  $10^{-1} \text{ M}$  of ZnO concentration which attained  $1.90 \times 10^{19} \text{ cm}^{-3}$ . Form this analyses, we can say that the doping of the PANI films with ZnO enhances the electrical conductivity ( $\sigma$ ) which is depends and proportional to the carrier concentration by the relation (2) [36]. Effectively the composite PANI/ZnO electropolymerized with  $10^{-1} \text{ M}$  concentration of ZnO in the electrolyte was predicted as the best conductive film.

$$\sigma = \mu q N_D \quad (2)$$

wher  $\sigma$  is electrical Conductivity,  $\mu$  is the Mobility of electrons,  $q$  is the elementary Charge of electron and  $N_D$  is the Concentration of carriers.



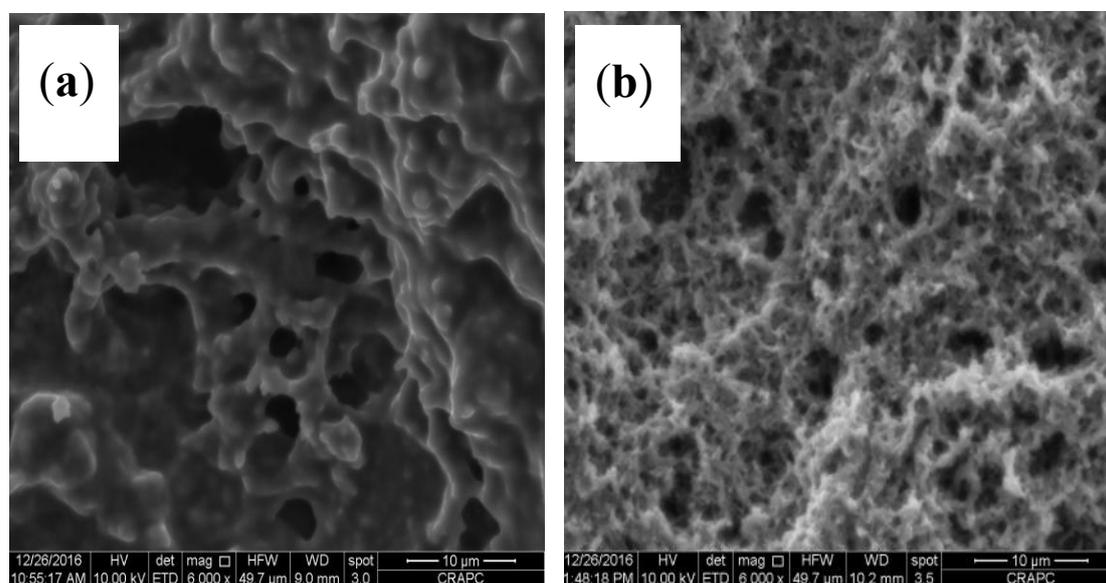
**Figure 3.** Mott schottky plot of PANI film and PANI/ZnO composite films electropolymerized at various concentrations of ZnO particles contact in  $1\text{M H}_2\text{SO}_4 + 0.1\text{M LiClO}_4$  with 500 Hz frequency.

**Table 1.** variation of the carrier concentration density, transmittance and band gap energies of PANI and PANI film and PANI/ZnO composite films with ZnO concentration.

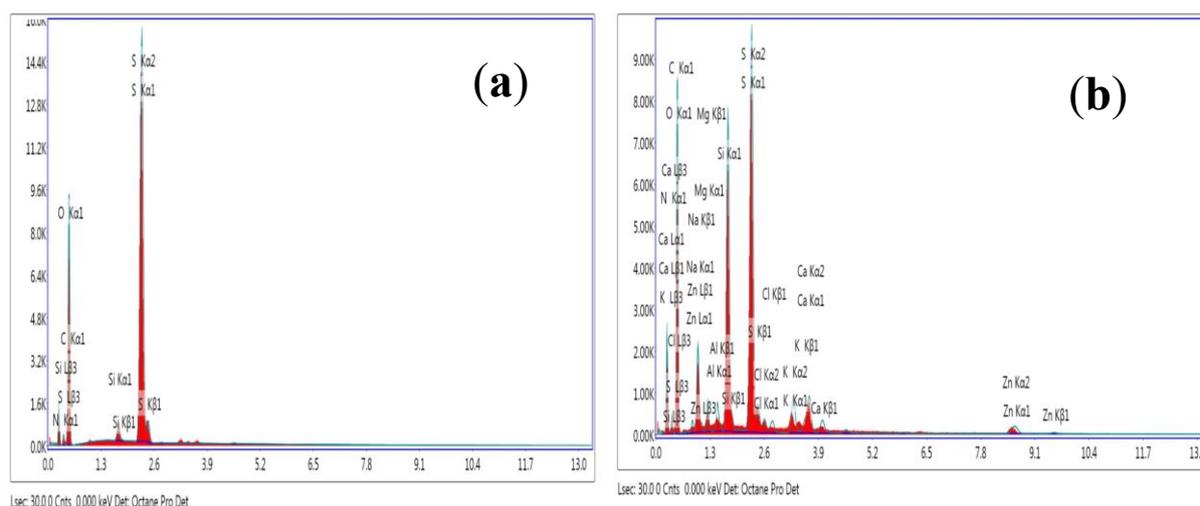
Parameters	PANI	PANI/ZnO	PANI/ZnO	PANI/ZnO
		$10^{-4}$ M	$10^{-2}$ M	$10^{-1}$ M
$N_D$ ( $\text{cm}^{-3}$ )	$2.87 \times 10^{17}$	$1.80 \times 10^{19}$	$7.06 \times 10^{18}$	$1.90 \times 10^{19}$
Transmittance (%)	70	18	84	17
$E_g$ (eV)	3.58	3.68	3.73	3.19

### 3.3. Scanning Electron Microscopy (SEM)

The surface morphology and elemental composition were studied through SEM coupled EDS analysis, Figure 4 presents the SEM micrographs of the PANI and PANI/ZnO composite films elaborate with  $10^{-1}$  M concentration of ZnO dopant in the electrolyte. From these images, it is clear that the morphology of the PANI was changed with presence of ZnO, the PANI texture was homogenous and porous and it becomes thinner with addition of the ZnO in the electrolyte. The corresponded EDS patterns of the samples (Figure 5) show several peaks, from the patterns (a) of the PANI film the carbon (C) and nitrogen (N) signals were characteristics of PANI polymer, soufre (S), oxygen (Cl) and chloride signals were going from sulfuric acid and  $\text{LiClO}_4$  dopants, respectively. The existed signal of Zn element in the patterns corresponded to the PANI/ZnO $10^{-1}$  M (patterns b) indicates the incorporation of the ZnO in the PANI matrix. Also the composition elements of ITO coated glass substrate were shown as peaks in the EDS patterns (In, Sn, O, Ca, Si, and Mg). Tables 2 and 3 show clearly the composition portion of these elements.



**Figure 4.** Scanning electron microscopy images: (a) PANI and (b) PANI/ZnO $10^{-1}$  M.



**Figure 5.** EDS patterns of (a) PANI, (b) PANI/ZnO  $10^{-1}$  M composite films.

**Table 2.** Composition elements corresponded to the PANI film.

Element	Weight %	Atomic %
C K	20.24	28.36
N K	5.91	7.11
O K	48.74	51.29
SiK	0.86	0.51
S K	24.25	12.73

**Table 3.** Composition elements corresponded to the PANI/ZnO  $10^{-1}$  M composite films.

Element	%Weight	%Atomic
C K	29.38	42.06
O K	40.97	44.04
MgK	0.27	0.19
SiK	3.30	2.02
S K	18.40	9.87
ClK	0.28	0.14
InL	2.45	0.37
SnL	0.75	0.11
CaK	0.60	0.26
ZnK	3.61	0.95

**Table 4.**

Parameters	PANI	PANI/ZnO $10^{-4}$ M	PANI/ZnO $10^{-2}$ M	PANI/ZnO $10^{-1}$ M
Transmittance (%)	70	18	84	17
Eg (eV)	3.58	3.68	3.73	3.19

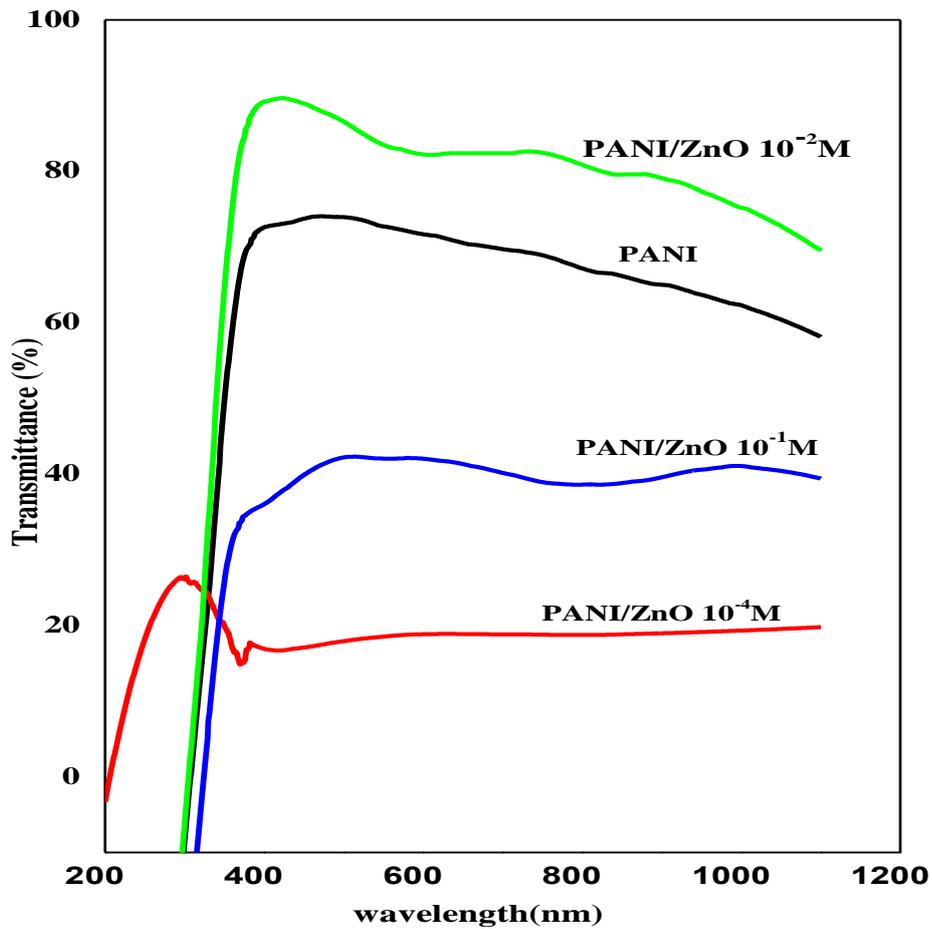
### 3.4. UV-VIS Study

Figure 6 shows the optical transmittance spectra of PANI and PANI/ZnO composite films electropolymerized at various concentrations of ZnO particles in the electrolyte, from this figure, all the prepared PANI and PANI/ZnO composite films exhibited an optical transmittance in the visible region with a sharp fundamental absorption edge observed in UV region at around 300 nm. It can be seen that the transmittance of the films decreases from 70 to the 18 and 17 with presence of ZnO in PANI matrix for the composite PANI/ZnO  $10^{-1}$  M and PANI/ZnO  $10^{-4}$  M, respectively. But the

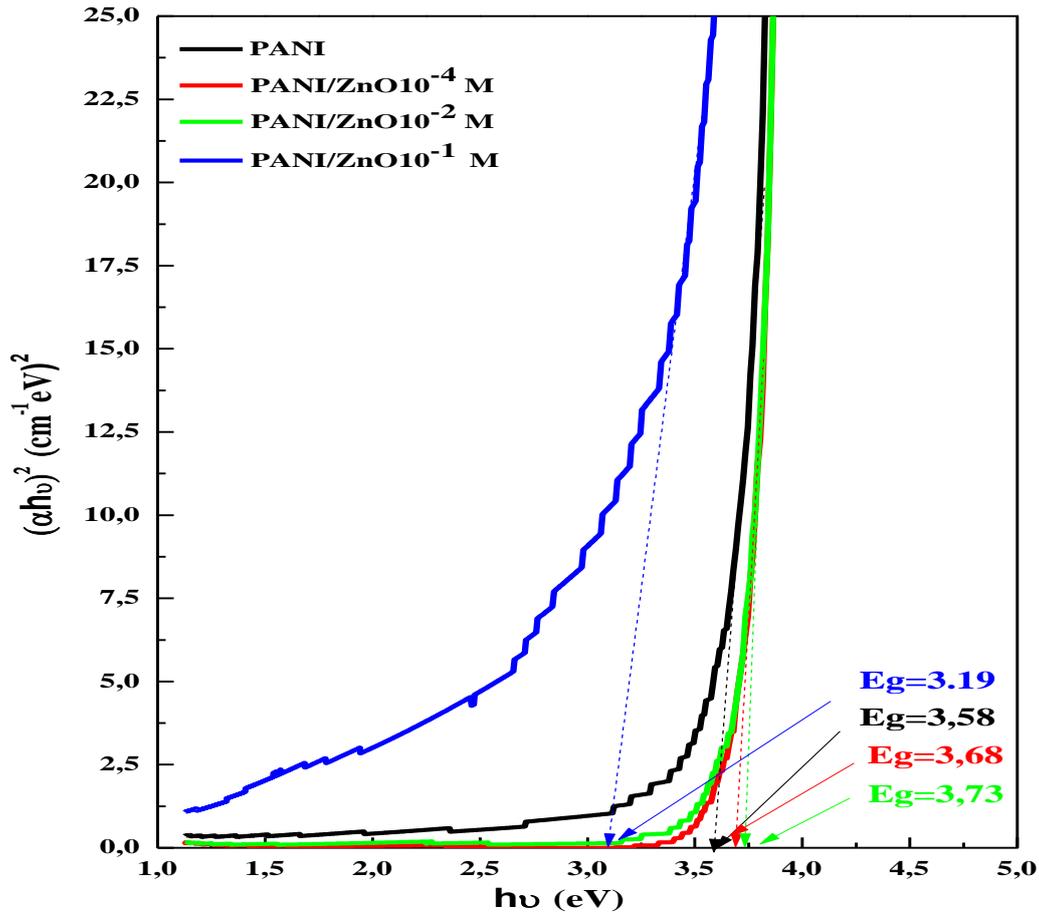
transmittance increases for PANI/ZnO  $10^{-2}$  M to attained 84%. The Tauc plots describe the dependence of the absorption coefficient and the photon energy for direct interband transitions [37] as the following equation

$$(\alpha h\nu)^2 = A (h\nu - E_g) \quad (3)$$

where A is a constant,  $\alpha$  is absorption coefficient, h Plank constant and  $\nu$  is the frequency of the radiation and  $E_g$  is the band gap energy. The optical band gap energies ( $E_g$ ) of PANI and PANI/ZnO composite films at various concentrations of ZnO particles were determined from the extrapolating of linear fit of  $(\alpha h\nu)^2$  versus  $h\nu$  at the  $\alpha h\nu = 0$  which shown in Figure 7. The Table 1 regroups the different values of transmittance and band gap energies corresponded to the PANI and PANI/ZnO films. It is observed that these optical characteristics were considerably influenced by the presence of different concentration of ZnO in the electrolyte.



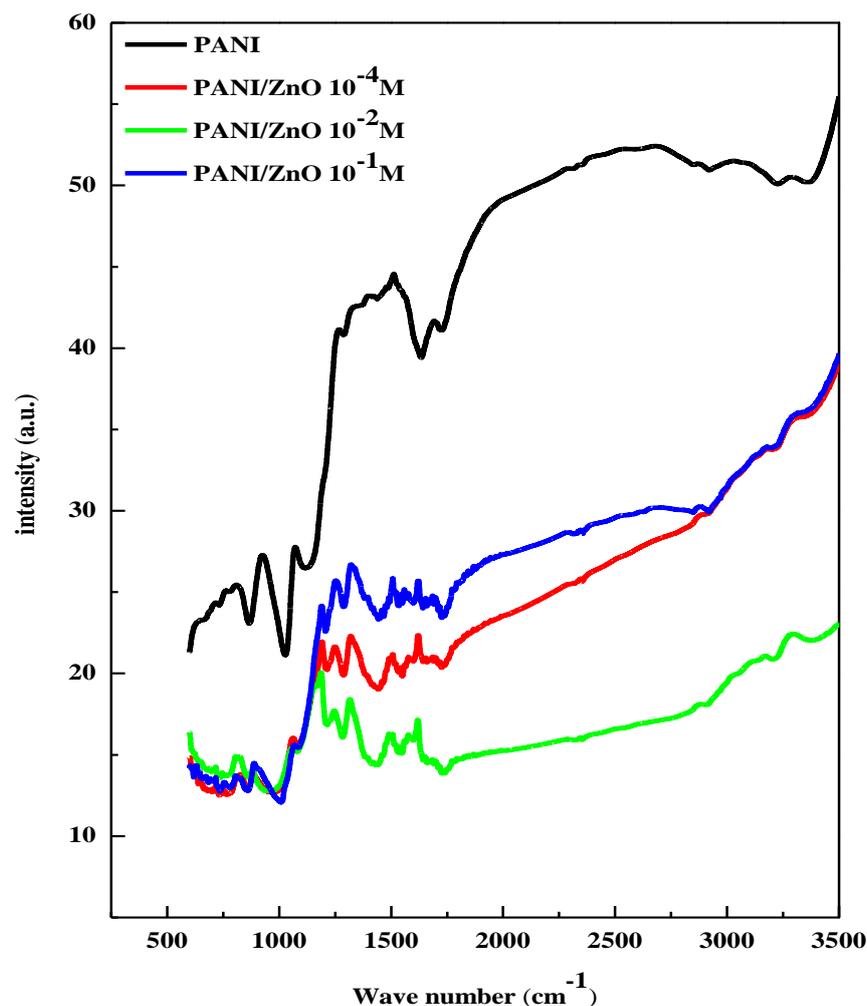
**Figure 6.** Transmittance spectra of PANI film and PANI/ZnO composite films electropolymerized at various concentrations of ZnO particles.



**Figure 7.** variation of  $(\alpha h\nu)^2$  as function of  $h\nu$  of PANI film and PANI/ZnO composite films electro polymerized at various concentrations of ZnO particles.

### 3.5. FTIR Characterization

Figure 8 shows the FTIR patterns of PANI and PANI/ZnO composite films electropolymerized with different concentrations of ZnO particles in the electrolyte. In the PANI patterns the absorption peaks at 1634 and 1439  $\text{cm}^{-1}$  characterized as the stretching mode of vibration corresponded to the C=C and C=N links of the quinoid and benzenoid units of the PANI. The low peak at 1287  $\text{cm}^{-1}$  was related to the C-N stretching mode of benzenoid circle and the IR absorption bands of angular deformation out of plan of C-H bond were found at 731 and 775  $\text{cm}^{-1}$  corresponded to the monosubstituted aromatic compounds presented in aniline monomer of PANI structure. For PANI/ZnO composite films with all concentration of ZnO particles, the same absorption peaks were found but with a good appearance of the peaks between 1200 to 1800 region and between  $\text{cm}^{-1}$  578 to 882  $\text{cm}^{-1}$  which can be attributed to the interaction between ZnO particle and PANI by formation bonding between ZnO and N-H group of PANI. Similar results were obtained for PANI/TiO<sub>2</sub>/ZnO composite by T. J. Brooms and al [38].



**Figure 8.** FTIR patterns of PANI and PANI/ZnO composite electropolymerized at different concentrations of ZnO particles.

#### 4. Conclusions

The polyaniline and polyaniline/ZnO composite films were elaborated by electrochemical polymerization using cyclic voltammetry technique on ITO substrate. Electrical properties were studied from the Mott-Schottky analysis, n-type conductivity was found for the PANI and the PANI/ZnO composite, the presence of ZnO doping in polyaniline matrix increases the carrier concentration which leading to ameliorate the electrical conductivity. The optimal concentration of ZnO particle in the electrolyte during electropolymerization was found  $10^{-1}$  M which was corresponded to highest carrier concentration of  $1.90 \times 10^{19} \text{ cm}^{-3}$ . The ZnO doping with  $10^{-4}$  and  $10^{-1}$  M concentration in the electrolyte has a negative effect on the optical transmittance of PANI but the ZnO doping with  $10^{-2}$  M ameliorate the transmittance value to the 82.30%. The FTIR analysis confirms the links presented in the polyaniline films and the interaction between ZnO and polyaniline for the polyaniline/ZnO composite films.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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