

Proceedings

# Study of Organic Radicals Generated upon Naphthoquinone-Hydantoins Reactions in Basic Aqueous Solution <sup>†</sup>

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**Abstract:** In this work, the synthesis of thiohydantoins from L-amino acids and isothiocyanates was studied, using a high purity method based on the Edman cycle. This process, expected to occur through Michael addition to a series of 1,4-naphthoquinone derivatives did not lead to the expected final product, rather leading to a colored mixture of compounds, where ESR spectra showed that some of them presented unpaired electrons, indicating that an electron transfer pathway was also involved in the reaction mixture. Detected organic radicals proved to be stable albeit the use of different experimental conditions. Voltammetric results indicate that the reactions lead to the formation of electroactive species, probably derived by homogeneous electron transfer between the reactive quinone moieties and the oxidizable urea functions within the hydantoin species.

**Keywords:** radicals; naphthoquinones; hydantoins; EPR

## 1. Introduction

Hydantoins (Figure 1) are a group of 5 membered heterocycles containing a reactive urea core (or an analog) which present multiple biological activities including antiviral, antiarrhythmic and antiandrogenic properties as well as being present in compounds used as fungicides, herbicides and pesticides [1–3]. These properties have made hydantoins an important element in fragment-based pharmaceuticals design [4].

Quinones can be chemically or electrochemically reduced to form hydroquinones, involving the formation of radical anion ( $Q^{\bullet-}$ ) and dianion species ( $Q^{2-}$ , Figure 2), which interact with polar biological elements like proteins, DNA and oxygen [5,6].

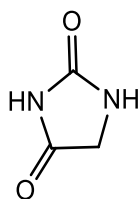


Figure 1. Hydantoin general structure.

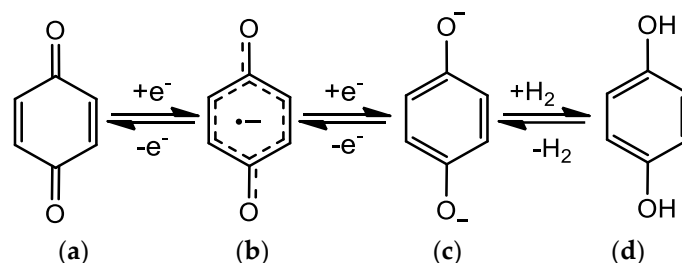


Figure 2. Quinone/Hydroquinone redox mechanism: (a) Quinone; (b) Anion Radical; (c) Dianion species; (d) Hydroquinone.

On the other hand, quinone compounds are well-known electrophilic compounds, enhanced by the electronic characteristics of the substituents and thus they are convenient Michael acceptors for 1,4 additions. In Michael addition processes, a nucleophilic addition to a  $\beta$  carbon of an  $\alpha,\beta$ -unsaturated carbonyl in the naphthoquinones can be done with substituted hydantoin as an electrophile. Acyclic amine additions to naphthoquinones are reported in the literature, where heptahydrate cerium trichloride is used as catalyst to add anilines to 5-hydroxy-1,4-naphthoquinones [7], and in more recent reports, the addition of different amino acids to 1,4-naphthoquinone using triethylamine and potassium hydroxide in dioxane-water by microwave irradiation has been described [8].

Interestingly, in the above mentioned reports, the formation of radical species was reported, which could be explained when considering hydantoin moieties as an electron source and thus oxidizable species. To obtain more information on this aspect of the reaction, further structure elucidation of the final products, including cyclic voltammetry analysis and Electron Spin Resonance (ESR) [9] was used and structure proposals of the obtained compounds are presented.

## 2. Methods

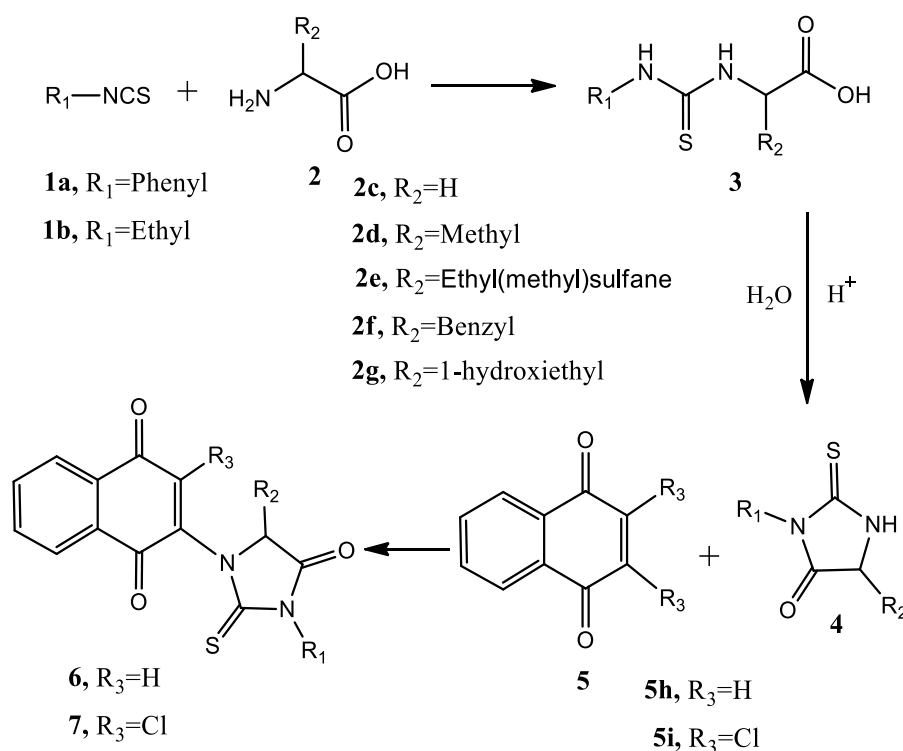
### 2.1. Thiohydantoins Synthesis

Thiohydantoins were synthesized using the Edman Cycle variation from Seung-Ju Yang et al. [10] with some modifications: A solution of isocyanate (phenylisocyanate, ethylisocyanate) in 1,4-dioxane:water 1:1 mixture at 0 °C was added to a proper L-Amino acid (1 eq.) and stirred for 15 min. Triethylamine (2 eq.) was added slowly to the solution and stirring continued for 1 h, followed by the slowly addition of concentrated HCl until pH was approximately 2; in these conditions, reaction continued up to 4 days. A saturated solution of  $\text{NaHCO}_3$  was used to adjust the solution pH = 6 and the formed precipitate was filtered and washed with a 1:1 1,4-dioxane:water mixture.

### 2.2. Naphthoquinone-Hydantoins Hybrids Synthesis

Commercially available 1,4-naphthoquinone, 2,3-dichloronaphthoquinone, hydantoin and thiohydantoin were used for the synthesis. A solution of 1 equivalent of hydantoin or thiohydantoin was prepared in 1:1 1,4-dioxane:water mixture in the presence of an alkali salt; the solution was introduced into the microwave oven and heated at 100 °C, 400 W for 5 min to deprotonate the acidic NH function. Afterwards, heating was stopped and 1 equivalent of either 1,4-naphthoquinone or 2,3-dichloronaphthoquinone was added to the solution and the reaction continued by the same

conditions for 20 min. Products were purified by column chromatography. An scheme of the process is shown in Figure 3.



**Figure 3.** General scheme of the employed synthetic procedure.

### 2.3. Products Characterization

$^1H$  and  $^{13}C$  Nuclear Magnetic Resonance (NMR) spectra were recorded on a Bruker Avance III 400 MHz spectrometer. High-Resolution Mass Spectra (HRMS) were measured with a Jeol JMS-AccuTOF through DART (Direct Analysis in Real Time). Infrared spectra were recorded on a Thermo Scientific NICOLET iS10 with ATR dispositive. Electron Spin Resonance (ESR) spectra were recorded on a Jeol JES-TE300 spectrometer. Melting points were determined using a Bicote-Stuart SMO 10 apparatus.

Cyclic voltammograms were obtained using a three electrode set-up, using a glassy carbon disk ( $\phi = 3$  mm) as working electrode, a platinum wire as counter electrode and a Ag|AgCl 0.1 M in  $CH_3CN$  was employed as reference electrode. Electrode potential values are referred to the Fc/Fc<sup>+</sup> couple as recommended by IUPAC. An AUTOLAB PGSTAT 302N potentiostat/galvanostat was employed for recording the electrochemical responses.

## 3. Results and Discussion

### 3.1. Thiohydantoin Synthesis

Thiohydantoin were synthesized with relatively good yields; even though the reaction takes a long time to proceed (more than 24 h), no further purification was needed (Table 1). All molecules were completely characterized by NMR, HRMS, IR and Melting Point determinations.

**Table 1.** Yields obtained in thiohydantoin synthesis.

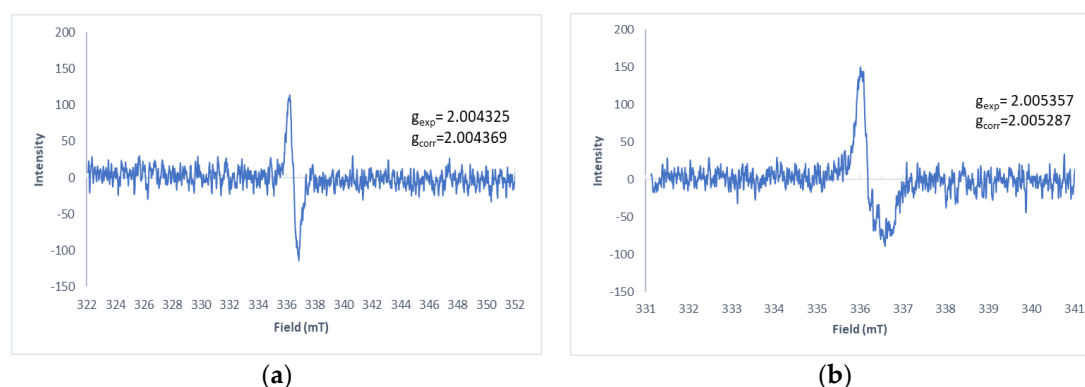
Compound <sup>1</sup>	Yield (%)
4ac	65
4ad	73
4ae	62
4af	69
4ag	58
4bd	55
4be	56
4bf	52

<sup>1</sup> First letter: isothiocyanate substituent, second letter: amino acid substituent.

### 3.2. Naphthoquinone-Hydantoin Hybrids Synthesis

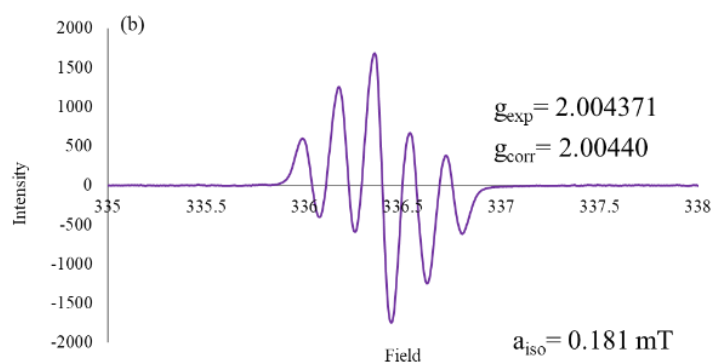
Thiohydantoin deprotonation with DABCO and Cs<sub>2</sub>CO<sub>3</sub> showed a color change from yellow (neutral) to pink (pH > 8), and then 2,3-dichloronaphthoquinone was added to the solution and the mixture were irradiated.

A pink/purple product formed for all the thiohydantoin, this product was extracted in DCM and purified by flash column chromatography (DCM:MeOH 9:1). NMR spectra were obtained for all the products but the proper assignment of the signals could not be performed even though different solvents were employed. As described above, it is widely known that hydantoin and quinones generate radicals as reaction intermediates, which would explain the difficulty to obtain good NMR spectra as the presence of a radical in the molecule overlaps due to the higher electron magnetic moment over the nuclear magnetic moment that can be up to 10<sup>3</sup> times larger [11]. Therefore, as an alternative, Electron Spin Resonance of the products was obtained. In these spectra, characteristic paramagnetic organic radicals were detected (Figure 4), however, the low amount of these species did not lead to proper resolution of the spectra.



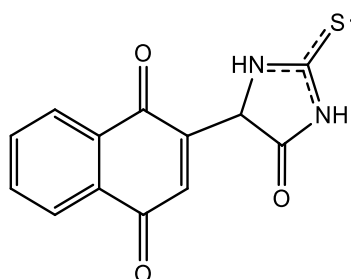
**Figure 4.** Solid state ESR spectra of (a) 7ad and (b) 7ac.

Experiments with commercial thiohydantoin lead us to obtain a well-defined spectrum in methanol (Figure 5). A quintuple with 1:2:3:2:1 relative intensities peaks was observed, this could be consistent with the presence of 2 nitrogen atoms within the structure, both with spin number of 1, according to  $2ni + 1$  formula for multiplicity or equivalent hydrogens from the aromatic ring, studies are now in progress to determine the possible structure.



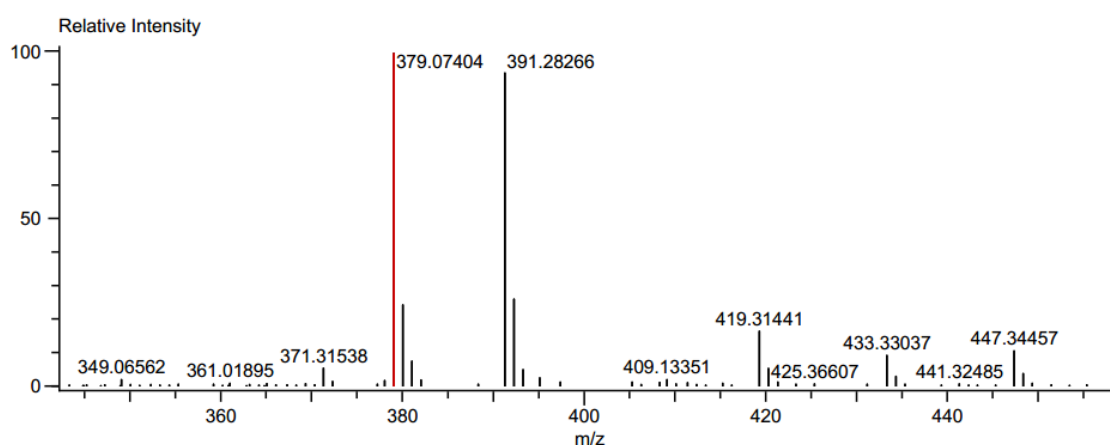
**Figure 5.** EPR spectrum for thiohydantoin-naphthoquinone hybrid.

Unsubstituted thiohydantoin possess 3 acidic hydrogens that can be subtracted to achieve the addition, so 3 possible products, including recombination, were expected. According to the quintuple obtained from the EPR spectra, a possible structure is one with the stabilization of the radical between the 2 nitrogen atoms, Figure 6



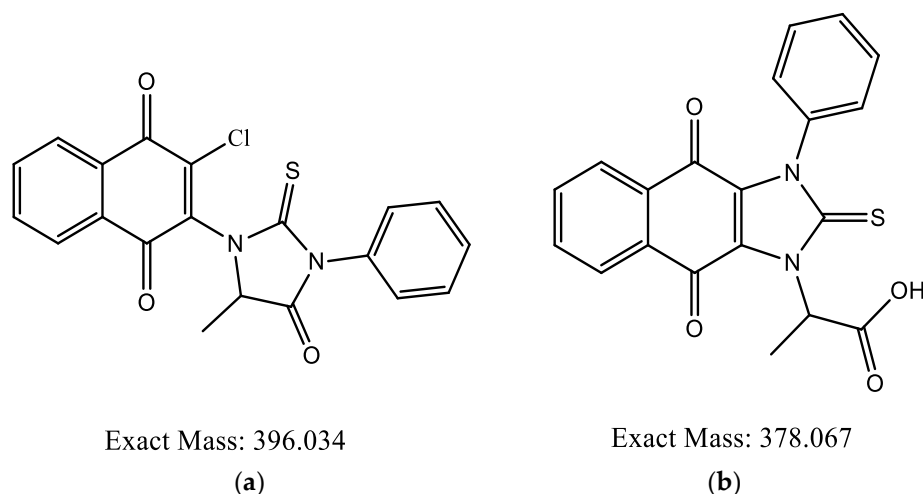
**Figure 6.** Proposed chemical structure for thiohydantoin-naphthoquinone hybrid.

Different synthetic pathways were explored to generate the thiohydantoin cycle over the naphthoquinone core by performing first the addition of amino acids to the naphthoquinones, followed by cyclization with isothiocyanates but this procedure lead to the same result. HRMS were also performed, small ion mass variations in the products were found. Product **6ad** had ion masses of 379.07404 and 391.28266  $m/z$  (Figure 7.), compared to the expected mass of 397.034  $m/z$  (Figure 8a).



Mass	Intensity	Calc. Mass	Mass Difference (mmu)	Mass Difference (ppm)	Possible Formula
379.07404	645817.83	379.07525	-1.21	-3.20	$^{12}\text{C}_{20}^{1}\text{H}_{15}^{14}\text{N}_2^{16}\text{O}_4^{32}\text{S}_1$

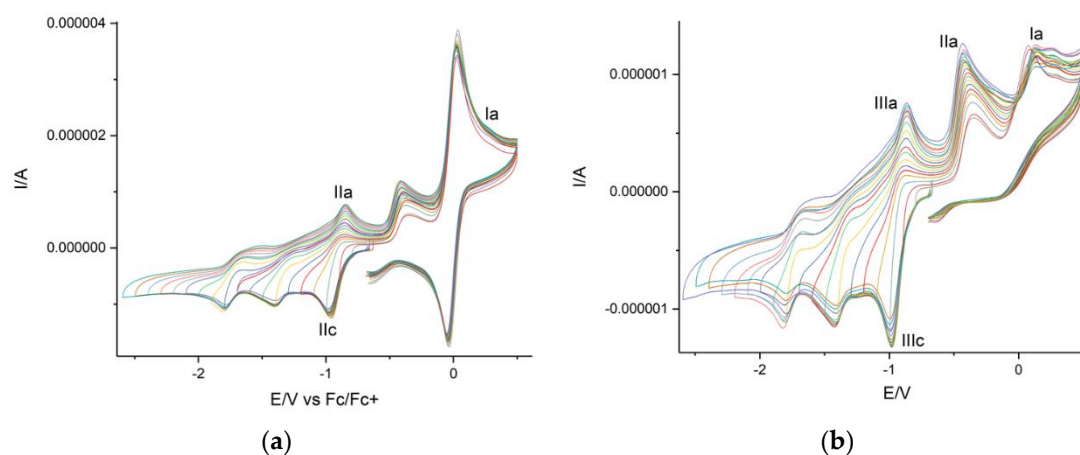
**Figure 7.** HRMS spectrum of product **6ad**.



**Figure 8.** Chemical structure for 6ad: (a) Expected; (b) Proposed.

According to literature, thiohydantoin exposed to alkaline solutions are easily ionized and in continuous exposition the cycle tends to open [12]. Therefore, it is possible that thiohydantoin could be added and later the cycle tends to open via the N-CO bond, cyclizing again by both nitrogen atoms forming the proposed chemical structure shown in Figure 8b. The same authors expressed that substituted thiohydantoin in C5 and N3 lead to a N1 atom remaining free, stabilizing their ionization between the free nitrogen and thiocarbonyl carbon [13], this reinforcing the results obtained from EPR spectra.

Cyclic voltammetry (CV) preliminary studies for obtained products showed the presence of anodic/cathodic reversible peaks corresponding to naphthoquinone reductions (Peaks IIc and IIIc in Figure 9) and the possible cathodic peak from thiohydantoin (Peak Ia in Figure 9b).



**Figure 9.** Cyclic Voltamperogram for thiohydantoin-naphthoquinone hybrid: (a) with ferrocene, (b) without ferrocene.

Michael additions of the thiohydantoin were done over  $\alpha,\beta$ -unsaturated carbonyl systems different from quinones under the same conditions and the expected products were formed and identified by NMR. Further studies of these compounds are already in process to obtain better information that allow us to explain correctly the chemical nature of these interesting products.

#### 4. Conclusions

Hydantoin ionization are potentiated by redox properties of naphthoquinones to form hydantoin-naphthoquinone hybrids that represent an interesting family of radical persistent compounds with yet uncertain chemical structure. Reaction conditions are useful for hydantoin

additions to  $\alpha,\beta$ -unsaturated carbonyl compound lacking of redox properties, this also represent an opportunity to keep exploring reactivity of hydantoins derivatives in different conditions. Future work is required in matter of ESR and CV to obtain the proper parameters for the radical structure elucidation, and it is in progress.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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